

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.

Remember: Percent yield is the ratio of actual yield to theoretical yield, multiplied by 100. Percent yield is always less than or equal to 100%.

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.

Remember: Percent yield is the ratio of actual yield to theoretical yield, multiplied by 100.



The reaction is balanced. The coefficients are 2, 1, and 2.

1. 10.0 g of H_2 and 10.0 g of O_2 are reacted. What is the limiting reagent? What is the percent yield?



2. 10.0 g of H_2 and 10.0 g of O_2 are reacted. What is the limiting reagent? What is the percent yield?

3. 10.0 g of H_2 and 10.0 g of O_2 are reacted. What is the limiting reagent? What is the percent yield?



4. 10.0 g of H_2 and 10.0 g of O_2 are reacted. What is the limiting reagent? What is the percent yield?



Percent Yield And Limiting Reactants Quiz

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