

The Periodic Law

SECTION 2

Select the correct answer for the following questions for the space provided.

1. _____ for the modern periodic table, elements are ordered
 - (a) according to increasing atomic mass
 - (b) according to Mendeleev's original design
 - (c) according to increasing atomic number
 - (d) based on how they were discovered
2. _____ Mendeleev stated that certain elements have similar chemical properties
 - (a) elements appeared in regular intervals and subsequent elements were arranged based on that spacing
 - (b) elements
 - (c) atomic numbers
 - (d) mass ratios
3. _____ The modern periodic law states that
 - (a) all new elements with the same properties are listed in the same group
 - (b) the physical and chemical properties of elements are functions of atomic number
 - (c) elements exhibit properties of both periods and groups
 - (d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot
4. _____ The discovery of the noble gases changed Mendeleev's periodic table
 - (a) adding a row
 - (b) adding a new period
 - (c) adding a new column
 - (d) adding a new group
5. _____ The noble gases that occupy the group 18 in the periodic table
 - (a) exhibit
 - (b) react
 - (c) are inert
 - (d) are gaseous
6. _____ Likewise, the first column in Group 1 has elements similar to H. The second column in the periodic table contains elements similar to
 - (a) He
 - (b) Li
 - (c) Be
 - (d) B
7. How groups in the periodic table are numbered?
 - (a) 1 to 18
 - (b) 1 to 10
 - (c) 1 to 16
 - (d) 1 to 18
8. How many periods are in the periodic table?
 - (a) 18
 - (b) 10
 - (c) 16
 - (d) 18
9. How many elements are in the periodic table?
 - (a) 118
 - (b) 108
 - (c) 116
 - (d) 118
10. What is the periodic law?
 - (a) The chemical and physical properties of elements are functions of their atomic number
 - (b) The chemical and physical properties of elements are functions of their atomic mass
 - (c) The chemical and physical properties of elements are functions of their atomic weight
 - (d) The chemical and physical properties of elements are functions of their atomic number

Periodic Law Chapter 5 Test Answers

Pieter Thyssen, Arnout Ceulemans

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author names.

Periodic Law Chapter 5 Test Answers:

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particles The trail follows the lines of parentage linking groups upstream to the unitary symmetry of the eightfold way of quarks and to the four dimensional symmetry of the hydrogen atom Along the way the theory of symmetry groups is gradually introduced with special emphasis on graphical representations The final challenge is to open up the structure of Mendeleev's table which goes beyond the symmetry of the hydrogen atom Breaking this symmetry to accommodate the multi electron atoms requires to leave the common ground of linear algebras and explore the potential of non linearity

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