

Unit 2: The Periodic Table Review Sheet

U2L1: PT History

- On the present day Periodic Table, elements are arranged in order of increasing atomic number
- **Groups:** vertical columns on the periodic table
 - Elements in the same group have similar chemical properties - because they have the same number of valence electrons
 - Valence electron: outermost electrons (lost # in the electron configuration)
- **Periods:** horizontal rows on the periodic table

U2L2: Parts of the Periodic Table

- Elements are classified as either metals, nonmetals, or metalloids.
 - Classification can be determined by location on the periodic table or by different properties

1

H

6

C

12.01

Atomic Mass

Symbol

Atomic Number

Electron Configuration

Selected Oxidation States

Commonly formed cations and anions

Notes: Numbers in parentheses are values for unstable or hypothetical elements or values in special cases.

18

He

Group 1

Li

Na

K

Rb

Cs

Fr

Group 2

Be

Mg

Ca

Sr

Ba

Ra

Group 3

Sc

Y

La

Ac

Group 4

Ti

Zr

Hf

Rf

Group 5

V

Nb

Ta

Db

Group 6

Cr

Mo

W

Sg

Group 7

Mn

Tc

Re

Bh

Group 8

Fe

Ru

Os

Hs

Group 9

Co

Rh

Ir

Mt

Group 10

Ni

Pd

Pt

Ds

Group 11

Cu

Ag

Au

Rg

Group 12

Zn

Cd

Hg

Cn

Group 13

B

Al

Ga

In

Tl

Uut

Group 14

C

Si

Ge

Sn

Pb

Uuq

Group 15

N

P

As

Sb

Bi

Uup

Group 16

O

S

Se

Te

Po

Uuh

Group 17

F

Cl

Br

I

At

Uus

Group 18

He

Ar

Kr

Xe

Rn

Uuo

Periodic Table Review Pt 2

Princeton Review



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