

# Periodicity - Electronegativity

Name: **Answer Key**

1. Fill in the blanks below, and then draw arrows on to the periodic table indicating the direction of the electronegativity trend.

Generally, electronegativity  
**increases**  
across periods and  
**increases**  
up groups.



2. List the elements below in order of increasing electronegativity. (Smallest to largest).

Ba, Ra, Ca, Sr, Mg  
**Ra, Ba, Ca, Mg, Sr**

Zn, Ni, Cu, Sc, Ti  
**Sc, Ti, Ni, Cu, Zn**

F, Br, Cl, Rb, Fr  
**Fr, Rb, Br, Cl, F**

Nb, Sn, Rb, Fe, I  
**Fr, Rb, Nb, Sn, I**

F, Li, Ba, K, Na  
**K, Na, Li, Ba, F**

P, Se, Ar, Mg, Ca  
**Ca, Mg, Se, P, Ar**

3. Are the elements below in order of decreasing electronegativity? If yes, write "yes". If not, write them in the correct order.

Ca, K, Sr, Br  
**Br, Sr, Ca, K**

S, Br, Cl, F  
**F, Cl, Br, S**

O, B, Li, Na  
**Yes**

Y, Sr, Rb, Fr  
**Yes**

Al, Sc, Se, Ba  
**Yes**

As, Pb, Te, Si  
**S, Te, Pb, As**

4. Briefly explain why electronegativity changes up groups.

Generally, electronegativity increases as you move up in a group. This is because the atomic number decreases, and as a result, there is less distance between the nucleus and valence electrons. In other words, there is a smaller atomic radius.

5. Briefly explain why electronegativity changes across periods.

As you move from left to right, across a period, the nuclear charge of elements increases faster than electron shielding. This results in an increased attraction to electrons.

# Periodicity Worksheet Answer Guide

**J Ma**



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