

- Reaction Rate = the change in concentration of reactant(s) or products per unit time. 12-3

↳ in the reaction on the last page (Table 12-1; Fig. 12-1) we have:



because being consumed (-)

$$\text{rate} = - \frac{\Delta[\text{NO}_2]}{\Delta t}$$

or

$$\text{rate} = \frac{\Delta[\text{NO}]}{\Delta t}$$

or

$$\text{rate} = 2 \left(\frac{\Delta[\text{O}_2]}{\Delta t} \right)$$



because only half as much produced. See the balanced equation.

* so, the reaction rate, or "rate", can be written in terms of any of the reactants or in terms of any of the products.

↳ To be consistent however, it's customary to represent the reaction rate in terms of the reactant(s), so:

$$\begin{array}{l} * \\ * \\ * \end{array} \boxed{\text{rate} = - \frac{\Delta[\text{NO}_2]}{\Delta t}}$$

- An Examination of Figure 12-1 on the Previous Page.

↳ at any single point on the NO_2 curve, we can calculate the instantaneous rate of the reaction.

↳ = calculated by finding the slope(m) of the line tangent to the curve at the specific time (t).

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