

The Periodic Law

Section 2

Answer the following questions for the space provided:

1. _____
 - (a) Mendeleev's periodic table, _____ is based
 - (i) according to the atomic weights.
 - (ii) according to Mendeleev's proposed change.
 - (iii) according to increasing atomic number.
 - (iv) based on the way they were discovered.
2. _____
 - (a) Mendeleev's periodic table _____ is called chemical properties of elements appeared in regular intervals whereas discovered later arranged elements in the way.
 - (i) atomic.
 - (ii) atomic number.
 - (iii) atomic mass.
 - (iv) atomic weight.
3. _____
 - (a) The modern periodic law states that
 - (i) all the elements with the same properties are placed in the same group.
 - (ii) the physical and chemical properties of elements _____ of _____.
 - (iii) elements exhibit _____ with periods and groups.
 - (iv) the chemical properties of elements can be grouped according to periodicity but physical properties _____.
4. _____
 - (a) The discovery of the noble gases changed Mendeleev's periodic table.
 - (i) adding a new _____.
 - (ii) period.
 - (iii) series.
 - (iv) group.
 - (v) block.
5. _____
 - (a) The noble gases that property of the noble gases is that they are
 - (i) inert.
 - (ii) unreactive.
 - (iii) radioactive.
 - (iv) the only _____.
6. _____
 - (a) List down the first six elements in Group 1. How are these elements classified? The second elements in this group have _____ number of
 - (i) 4.
 - (ii) 11.
 - (iii) 10.
 - (iv) 12.
7. The groups in Mendeleev's periodic table are numbered (a) 1 to 10.
 - (i) How many _____ (b) How many _____
 - (ii) How many _____ are in _____ group?
 - (iii) How many _____ are there in _____ group?
8. How many _____ are there in _____ group?
 - (i) How many _____ are there in _____ group?
 - (ii) How many _____ are there in _____ group?
 - (iii) How many _____ are there in _____ group?

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