

## Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. The charge on an ion is  
a. always positive.  
b. always negative.  
c. either positive or negative.  
d. zero.
- \_\_\_\_\_ 2. According to the octet rule, a calcium atom has a tendency to  
a. lose one electron.  
b. lose two electrons.  
c. gain one electron.  
d. gain two electrons.
- \_\_\_\_\_ 3. If a compound forms by ionic bonding, which is *not* true?  
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.  
b. The net charge of the compound is zero.  
c. The compound contains just two atoms, each of opposite charge.  
d. Several ions group together in a tightly packed structure.
- \_\_\_\_\_ 4. The only property listed that is *not* characteristic of ionic compounds is  
a. high melting point.  
b. hardness.  
c. lack of crystal structure.  
d. good conductor of electricity.
- \_\_\_\_\_ 5. Which formula listed below represents a polyatomic ion?  
a.  $\text{NH}_4^+$   
b.  $\text{H}_2\text{SO}_4$   
c.  $\text{NaCl}$   
d.  $\text{H}_2\text{O}$
- \_\_\_\_\_ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because  
a. ionic substances tend to vaporize at room temperature.  
b. ionic substances are brittle.  
c. attractive forces between ions are greater than the attractive forces between molecules.  
d. the numbers of positive and negative charges are equal in an ionic compound.

# Section 6 1 Review Chemical Bonding Answer Key

**William H. Brown, Thomas Poon**



## **Section 6 1 Review Chemical Bonding Answer Key:**

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## **Table of Contents Section 6 1 Review Chemical Bonding Answer Key**

1. Understanding the eBook Section 6 1 Review Chemical Bonding Answer Key
  - The Rise of Digital Reading Section 6 1 Review Chemical Bonding Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Section 6 1 Review Chemical Bonding Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Section 6 1 Review Chemical Bonding Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Section 6 1 Review Chemical Bonding Answer Key

- Personalized Recommendations
- Section 6 1 Review Chemical Bonding Answer Key User Reviews and Ratings
- Section 6 1 Review Chemical Bonding Answer Key and Bestseller Lists
- 5. Accessing Section 6 1 Review Chemical Bonding Answer Key Free and Paid eBooks
  - Section 6 1 Review Chemical Bonding Answer Key Public Domain eBooks
  - Section 6 1 Review Chemical Bonding Answer Key eBook Subscription Services
  - Section 6 1 Review Chemical Bonding Answer Key Budget-Friendly Options
- 6. Navigating Section 6 1 Review Chemical Bonding Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Section 6 1 Review Chemical Bonding Answer Key Compatibility with Devices
  - Section 6 1 Review Chemical Bonding Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Section 6 1 Review Chemical Bonding Answer Key
  - Highlighting and Note-Taking Section 6 1 Review Chemical Bonding Answer Key
  - Interactive Elements Section 6 1 Review Chemical Bonding Answer Key
- 8. Staying Engaged with Section 6 1 Review Chemical Bonding Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Section 6 1 Review Chemical Bonding Answer Key
- 9. Balancing eBooks and Physical Books Section 6 1 Review Chemical Bonding Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Section 6 1 Review Chemical Bonding Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Section 6 1 Review Chemical Bonding Answer Key
  - Setting Reading Goals Section 6 1 Review Chemical Bonding Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 6 1 Review Chemical Bonding Answer Key

- Fact-Checking eBook Content of Section 6 1 Review Chemical Bonding Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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