

REVIEW: Chapter 8 - Covalent Bonding

- Typically, how many bonds & lone pairs will each have when covalently bonding?
 - Boron – 3 bonds; 0 lone pairs
 - Silicone – 4 bonds; 0 lone pairs
 - Neon – 0 bonds; 4 lone pairs
 - Iodine – 1 bond; 3 lone pairs
 - Nitrogen – 3 bonds; 1 lone pair
 - Oxygen – 2 bonds; 2 lone pairs
- Which elements do not follow the octet rule? (Hint – there are four we went over)
 - H & B never follow the octet rule.
 - P & S sometimes don't follow the octet rule.
- What type of bond will each combination of atoms form?
 - F & N – 2 nonmetals, therefore covalent
 - Al & O – 1 metal/1 nonm., therefore ionic
 - Na & S – 1 metal/1 nonm., therefore ionic
 - B & Cl – 2 nonmetals, therefore covalent
 - Ca & I – 1 metal/1 nonm., therefore ionic
 - Fe & Cu – 2 metals, therefore metallic
- What is the difference between:
 - Molecules & Compounds?
Compounds can be made using all types of bonds (ionic, covalent or metallic), but molecules can only be made using covalent bonds
 - Monatomic elements & Diatomic molecules?
Monatomic elements exist as single atoms; diatomic molecules are made up of two of the same atom
 - Structural formulas, Empirical formulas, & Molecular formulas?
All three are for covalent bonds only! Molecular show the actual number of atoms in the molecule; Empirical give the lowest whole number ratio; Structural formulas are not really formulas, they are drawing of the molecule.
 - Polar bonds & Non-polar bonds?
Both describe covalent bonds! In a polar bond the electrons are NOT shared equally; in a non-polar bond electrons are shared equally
- Classify each of the following as an atom or a molecule:
 - Be, atom
 - CO, molec.
 - N₂, molec.
 - H₂O, molec.
 - Ne, atom

Section 8 Covalent Bonding Workbook Answers

L Reisser



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