#### Balancing Chemical Equations Worksheet

1. 
$$2 H_2 + 0_2 \rightarrow 2 H_2O$$

3. 
$$S_8 + 12 O_2 \rightarrow 8 SO_3$$

4. 
$$2 N_2 + 0_2 \rightarrow 2 N_20$$

5. 
$$2 \text{ HgO} \rightarrow 2 \text{ Hg} + 0$$

6. 
$$6 \, CO_2 + 6 \, H_2O \Rightarrow C_6H_{12}O_6 + 6 \, O_2$$

7. 
$$Zn + Z HCI \rightarrow ZnCl_2 + H_2$$

9. 
$$\underline{2}$$
 Na +  $\underline{2}$  H<sub>2</sub>O  $\Rightarrow$   $\underline{2}$  NaOH +  $\underline{\phantom{0}}$  H<sub>2</sub>

12. \_\_\_\_ 
$$CO_2$$
 + \_\_\_  $Z$   $NH_3  $\rightarrow$  ____  $OC(NH_2)_2$  + ____  $H_2O$$ 

13. 
$$_{-}^{4}$$
 Si<sub>2</sub>H<sub>3</sub> +  $_{-}^{7}$  O<sub>2</sub>  $\rightarrow$   $_{-}^{8}$  SiO<sub>2</sub> +  $_{-}^{6}$  H<sub>2</sub>O<sub>3</sub>

14. 
$$2 \text{ Al}(OH)_3 + 3 \text{ H}_2SO_4 \rightarrow Al_2(SO_4)_3 + 6 \text{ H}_2O$$

17. 
$$2 C_7H_6O_2 + 15 O_2 \rightarrow 14 CO_2 + 5 H_2O$$

19. \_\_\_\_\_ FeS<sub>2</sub> + \_\_\_\_\_ O<sub>2</sub> 
$$\Rightarrow$$
 \_\_\_\_\_ Fe<sub>2</sub>O<sub>3</sub> + \_\_\_\_\_ SO<sub>2</sub>

21. \_\_\_\_ Fe<sub>2</sub>O<sub>3</sub> + \_\_3\_ H<sub>2</sub> 
$$\rightarrow$$
 \_\_2\_ Fe + \_\_3\_ H<sub>2</sub>O

22. \_\_\_\_ Na<sub>2</sub>CO<sub>3</sub> + \_\_2\_ HCl 
$$\rightarrow$$
 \_2\_ NaCl + \_\_\_\_ H<sub>2</sub>O + \_\_\_ CO<sub>2</sub>

24. \_\_\_\_C<sub>7</sub>H<sub>16</sub> + \_//\_O<sub>2</sub> 
$$\rightarrow$$
 \_7 CO<sub>2</sub> + \_8 H<sub>2</sub>O

25. 
$$P_4 + 5 O_2 \rightarrow 2 P_2O_5$$

# **Section Reactions And Equations Answer Key**

**Sabine Zange** 

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