

## Matter—Properties and Changes

### Section 3.1 Properties of Matter

pages 70–75

#### Problem-Solving Lab

1. Explain why the flow of a compressed gas must be controlled for practical and safe use.

The flow of compressed gas must be controlled to control the amount and the rate at which gas is released.

2. Predict what would happen if the valve on a full tank of compressed gas were suddenly opened all the way or if the tank were accidentally punctured.

Without the regulator device, the gas would rush out of the tank with a force powerful enough to transform the tank into a dangerous, uncontrolled projectile.

### Section 3.1 Assessment

page 75

1. Create a table that describes the three common states of matter in terms of their shape, volume, and compressibility.

	Volume	Shape	Compressibility
Solid	Definite	Definite	Incompressible
Liquid	Definite	Takes shape of container and fills container to the extent of its own volume	Virtually incompressible
Gas	Fills volume of container	Takes shape of container	Compressible

2. Describe the characteristics that identify a sample of matter as being a substance.

The sample of matter must have a uniform and unchanging composition to be a substance.

3. Classify each of the following as a physical or chemical property.
  - a. Iron and oxygen form rust.  
chemical

- b. Iron is more dense than aluminum.  
physical

- c. Magnesium burns brightly when ignited.  
chemical

- d. Oil and water do not mix.  
physical

- e. Mercury melts at  $-39^{\circ}\text{C}$ .  
physical

4. **Organize** Create a chart that compares physical and chemical properties. Give two examples for each type of property.

The chart should make clear that physical properties can be observed without changing the composition of the sample, which is not the case for chemical properties. Mass and density are examples of physical properties. Fermentation and rusting are examples of chemical properties.

### Section 3.2 Changes in Matter

pages 76–79

#### Practice Problems

page 78

5. Use the data in the table to answer the following questions.

Aluminum and Liquid Bromine Reaction		
	Before Reaction	After Reaction
Aluminum	10.3 g	0.0 g
Liquid bromine	100.0 g	8.5 g
Compound	0.0 g	

How many grams of bromine reacted? How many grams of compound were formed?

$$\text{amount of bromine that reacted} = 100.0 \text{ g} - 8.5 \text{ g} = 91.5 \text{ g}$$

$$\text{amount of compound formed} = 100.0 \text{ g} + 10.3 \text{ g} - 8.5 \text{ g} = 101.8 \text{ g}$$

# Section Reinforcement Chemical Changes Answers

**Kenneth L. Henold, Frank Walmsley**



## **Section Reinforcement Chemical Changes Answers:**

*Physical and Chemical Changes* Edward P. Ortleb, Richard Cadice, 1993-09-01 General chemistry information including everything from matter to radioactivity For grades 5 to 9

**Chemistry: Molecules, Matter, and Change Media Activities Book** Loretta Jones, Carl Hoeger, Peter William Atkins, Regina Schoenfield-Tacher, 2000-01-15 The Media Activity Book MAB for Jones Atkins Chemistry molecules matter and change contains chapters with lists and descriptions of some of the media available as you study the chapter Each activity begin with a specific textbook reference Then you are given a time estimate of how long it will take to use the media An M media icon in the margin of the textbook means that media exists to support that area of text The media is found in three different places on the website and on two CDs

Physical Science Anted, Grades 6-9 David W. Appenbrink, 1986

**Holt Science and Technology** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001-07

*Physical Forces & Chemical Changes* Wild Goose Company, 2002 Solve the mysteries behind amazing forces at work all around us Examine the properties of water It s a solid No it s a liquid No wait it s a gas Use a coffee filter to examine the science of color Clean pennies chemically build a statue with nails and make balloon rockets

**Failure, Distress and Repair of Concrete Structures** Norb Delatte, 2009-10-26 Understanding and recognising failure mechanisms in concrete is a fundamental pre requisite to determining the type of repair or whether a repair is feasible This title provides a review of concrete deterioration and damage as well as looking at the problem of defects in concrete It also discusses condition assessment and repair techniques Part one discusses failure mechanisms in concrete and covers topics such as causes and mechanisms of deterioration in reinforced concrete types of damage in concrete structures types and causes of cracking and condition assessment of concrete structures Part two reviews the repair of concrete structures with coverage of themes such as standards and guidelines for repairing concrete structures methods of crack repair repair materials bonded concrete overlays repairing and retrofitting concrete structures with fiber reinforced polymers patching deteriorated concrete structures and durability of repaired concrete With its distinguished editor and international team of contributors Failure and repair of concrete structures is a standard reference for civil engineers architects and anyone working in the construction sector as well as those concerned with ensuring the safety of concrete structures Provides a review of concrete deterioration and damage Discusses condition assessment and repair techniques standards and guidelines

**Chemical Principles, Properties, and Reactions** Kenneth L. Henold, Frank Walmsley, 1984

**International Marine Engineering**, 1918

Marine Engineering Log, 1918

*Limit Analysis and Concrete Plasticity, Second Edition* M.P. Nielsen, 1998-12-29 Limit Analysis and Concrete Plasticity Second Edition covers the most relevant topics related to plastic design methods providing a reliable and superior alternative to existing empirical methods Fully updated and containing more extensive coverage this second edition includes numerical methods and computer code for solving problems incorporating methods into Eurocode 2 the common concrete standard for the whole of Europe This edition Emphasizes

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**Journal of Medicine and Surgery** ,1909 [CBSE Class 12 Chemistry English Story Chapter 8](#) Priti Singhal,2024-11-11

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