

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

- 1. Represent the processes of a reaction.
- 2. Explain how to write a balanced equation.
- 3. List the steps for writing a complete ionic equation.

Vocabulary

- 1. chemical equation
- 2. balanced equation
- 3. catalyst
- 4. coefficient
- 5. skeletal equation

Part A Completion

Write the coefficient for each in the balanced equation, and balance the equation by changing the coefficient for each substance. Write balanced ionic equations for each reaction.

1. A solid compound is heated, producing a gas and a liquid. _____
 The products themselves are heated, changing into _____, but then
 become homogeneous liquids. Write the _____ equation and write
 the net ionic equation for _____, a chemical reaction between two
 homogeneous liquids. See page _____ in your textbook for
 instructions for the proper format of both of the _____, since neither is
 the liquid. The equation was written before the heating. _____
 The above reaction is reversible.

2. A substance is heated, and the product is a solid. The substance is _____, but a solid _____, and the gas
 _____, a substance that is a solid. _____, _____, _____, _____
 need to be added to the equation to balance the equation. Write the balanced
 equation.

- 1. _____
- 2. _____
- 3. _____
- 4. _____
- 5. _____
- 6. _____
- 7. _____
- 8. _____
- 9. _____
- 10. _____
- 11. _____
- 12. _____
- 13. _____
- 14. _____
- 15. _____
- 16. _____
- 17. _____
- 18. _____
- 19. _____
- 20. _____

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