

Colligative properties are properties that depend only upon the number of solute atoms, ions, or molecules in a solution and not on the nature of those atoms, ions or molecules. Freezing point depression and boiling point elevation are examples of colligative properties. Raoult discovered that the addition of solute particles causes the boiling point of a solution to be elevated and the freezing point to be depressed. It has been found that 1 mole of nonvolatile, nonionizing solute particles will raise the boiling point of 1.0 kg of water by 0.52 °C, to 100.52°C at 1 atm of pressure. The same concentration will lower the freezing point of 1.0 kg of water by 1.86°C, to -1.86°C. These two numbers apply only to water and are called the molal boiling point and freezing point constants for water. Other solvents have different constants.

PROPERTY	BOILING POINT ELEVATION	FREEZING POINT DEPRESSION
EQUATION	Boiling Point Elevation: A solution will boil at a higher temperature than the pure solvent. $\Delta T_b = i K_b m$ ΔT_b = change in boiling temp. in (°C) <i>i</i> = a unitless constant (called the van't Hoff factor) associated with the degree of dissociation of the solute in a solvent. K_b = molal boiling point elevation constant (°C/m) For H₂O: K_b = .52°C/m <i>m</i> = molality of the solution	Freezing Point Depression: A solution will freeze at a lower temperature than the pure solvent. $\Delta T_f = i K_f m$ ΔT_f = change in freezing temp. in (°C) <i>i</i> = a unitless constant (called the van't Hoff factor) associated with the degree of dissociation of the solute in a solvent. K_f = molal freezing point lowering constant (°C/m) For H₂O: K_f = 1.86°C/m <i>m</i> = molality of the solution
	<i>i</i> = 1 for substances that do not ionize 2 for substances which ionize into two ions like LiCl 3 for substances which ionize into three ions like CaCl ₂ ...etc. <i>i</i> may have to be calculated if a % of dissociation is given for a solute. If the solvent is not water you must find its K_b .	<i>i</i> = 1 for substances that do not ionize 2 for substances which ionize into two ions like LiCl 3 for substances which ionize into three ions like CaCl ₂ ...etc. <i>i</i> may have to be calculated if a % of dissociation is given for a solute. If the solvent is not water you must find its K_f .

EXAMPLE #1

If 90.0 g of nonionizing C₆H₁₂O₆, are dissolved in 255 g of H₂O, what will be the boiling point of the resulting solution?

First calculate the molality (*m*) of the solution:

$$m = \frac{92\text{g}}{\left(180 \frac{\text{g}}{\text{mole}}\right)(0.255 \text{ kg of solvent})} = 2.0 \text{ m}$$

Next apply the boiling point elevation equation:

$$[\textit{i} = 1 \text{ for a nonionizing solute.}]$$

$$\Delta T_b = i K_b m$$

$$\Delta T_b = (1)(0.52^\circ\text{C/m})(2.0 \text{ m})$$

$$\Delta T_b = 1.04^\circ\text{C}$$

$$\text{BP} = (\Delta T_b + 100.00^\circ\text{C})$$

$$\text{BP} = (1.04^\circ\text{C} + 100.00^\circ\text{C}) = \underline{101.04^\circ\text{C}}$$

Section Review 16 3 Colligative Properties Of Solutions

**David Shubert, David C Shubert, John
Leyba**



Section Review 16 3 Colligative Properties Of Solutions:

Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 *Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today* Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application **Foundations of College Chemistry** Morris Hein, Susan Arena, Cary Willard, 2023 Foundations of College Chemistry 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need This International Adaptation offers new and updated content with improved presentation of all course material It builds on the strengths of previous editions including clear explanations and step by step problem solving The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry There is new and expanded coverage of polarizing power and polarizability Fajans rules collision number and mean free path abnormal molecular masses and van t Hoff factor and applications of radioactivity **5 Steps to a 5: AP Chemistry 2021** Mary Millhollon, Richard H. Langley, 2020-10-02 MATCHES THE LATEST EXAM In this hybrid year let us supplement your AP classroom experience with this multi platform study guide The immensely popular 5 Steps to a 5 AP Chemistry guide has been updated for the 2020 21 school year and now contains 3 full length practice exams available both in the book and online that reflect the latest exam Up to Date Resources for COVID 19 Exam Disruption Access to a robust online platform Comprehensive overview of the AP Chemistry exam format Hundreds of practice exercises with thorough answer explanations Proven strategies specific to each section of the test A self guided study plan including flashcards games and more online *MCAT General Chemistry Review* Alexander Stone Macnow, 2016-07-05 The most efficient learning for the MCAT results you want Kaplan s MCAT General Chemistry Review has all the information and strategies you need to score higher on the MCAT This book features more practice than any other guide plus targeted subject review questions opportunities for self analysis a complete online center and thorough instruction on all of the general chemistry concepts necessary for MCAT success from the creators of the 1 MCAT prep course page 4 of cover [Class 8-12 Chemistry Questions and Answers PDF](#) Arshad Iqbal, The Class 8 12 Chemistry Quiz Questions and Answers PDF Grade 8 12 Chemistry Competitive Exam Questions Chapter 1 15 Practice Tests Chemistry Textbook Questions for Beginners includes Questions to solve problems with hundreds of class questions Class 8 12 Chemistry Questions and Answers PDF book covers basic

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Edition David Shubert, David C Shubert, John Leyba, Sharon Niemann, 2017-01-25 Promotes ease of understanding with a unique problem solving method and new clinical application scenarios With a focus on chemistry and physics content that is directly relevant to the practice of anesthesia this text delivers in an engaging conversational style the breadth of scientific information required for the combined chemistry and physics course for nurse anesthesia students Now in its third edition the text is updated and reorganized to facilitate a greater ease and depth of understanding It includes additional clinical application scenarios detailed step by step solutions to problems and a Solutions Manual demonstrating a unique method for solving chemistry and physics problems and explaining how to use a calculator The addition of a third author a practicing nurse anesthetist provides additional clinical relevance to the scientific information Also included is a comprehensive listing of need to know equations The third edition retains the many outstanding learning features from earlier editions including a special focus on gases the use of illustrations to demonstrate how scientific concepts relate directly to their clinical application in anesthesia and end of chapter summaries and review questions to facilitate self assessment Ten on line videos enhance teaching and learning and abundant clinical application scenarios help reinforce scientific principles and relate them to day to day anesthesia procedures This clear easy to read text will help even the most chemistry and physics phobic students to master the foundations of these sciences and competently apply them in a variety of clinical situations New to the Third Edition The addition of a third co author a practicing nurse anesthetist provides additional clinical relevance Revised and updated to foster ease of understanding Detailed step by step solutions to end of chapter problems Solutions Manual providing guidance on general problem solving calculator use and a unique step by step problem solving method Additional clinical application scenarios Comprehensive list of all key equations with explanation of symbols New instructor materials include PowerPoint slides Updated information on the gas laws Key Features Written in an engaging conversational style for ease of understanding Focuses solely on chemistry and physics principles relevant to nurse anesthetists Provides end of chapter summaries and review questions Includes abundant illustrations highlighting application of theory to practice

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