

Unit 2 – Electrons and Periodic Behaviour

2.1

1.

Models of the Atom.

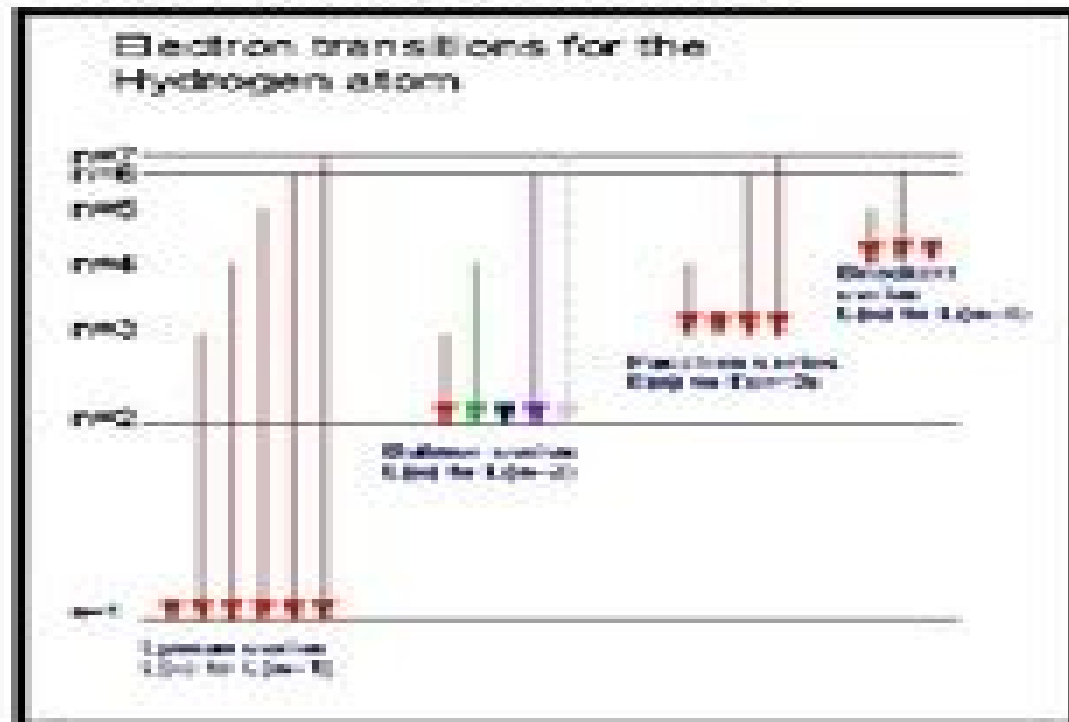
The Bohr Model of the Atom

A. Electrons Orbit, or Energy Levels

1. Electrons can orbit the nucleus only in allowed paths or orbits
2. The energy of the electrons is greater when it is to orbit further from the nucleus
3. The atom achieves the ground state when it is in the closest possible position around the nucleus
4. Electromagnetic radiation is emitted when electrons move closer to the nucleus

B. Energy transitions

1. Electrons of atoms are fixed and definite quantities
2. Energy transitions occur in jumps of discrete amounts of energy
3. Electrons only lose energy when they move to a lower energy state



C. Drawings of the Bohr Model

1. Doesn't work for atoms larger than hydrogen (more than one electron)
2. Doesn't explain chemical behaviour

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Table of Contents Section Tour Of The Periodic Table Holt Chemistry Answers

1. Understanding the eBook Section Tour Of The Periodic Table Holt Chemistry Answers
 - The Rise of Digital Reading Section Tour Of The Periodic Table Holt Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Section Tour Of The Periodic Table Holt Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section Tour Of The Periodic Table Holt Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section Tour Of The Periodic Table Holt Chemistry Answers
 - Personalized Recommendations
 - Section Tour Of The Periodic Table Holt Chemistry Answers User Reviews and Ratings
 - Section Tour Of The Periodic Table Holt Chemistry Answers and Bestseller Lists
5. Accessing Section Tour Of The Periodic Table Holt Chemistry Answers Free and Paid eBooks
 - Section Tour Of The Periodic Table Holt Chemistry Answers Public Domain eBooks
 - Section Tour Of The Periodic Table Holt Chemistry Answers eBook Subscription Services
 - Section Tour Of The Periodic Table Holt Chemistry Answers Budget-Friendly Options

6. Navigating Section Tour Of The Periodic Table Holt Chemistry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Section Tour Of The Periodic Table Holt Chemistry Answers Compatibility with Devices
 - Section Tour Of The Periodic Table Holt Chemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section Tour Of The Periodic Table Holt Chemistry Answers
 - Highlighting and Note-Taking Section Tour Of The Periodic Table Holt Chemistry Answers
 - Interactive Elements Section Tour Of The Periodic Table Holt Chemistry Answers
8. Staying Engaged with Section Tour Of The Periodic Table Holt Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section Tour Of The Periodic Table Holt Chemistry Answers
9. Balancing eBooks and Physical Books Section Tour Of The Periodic Table Holt Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section Tour Of The Periodic Table Holt Chemistry Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Section Tour Of The Periodic Table Holt Chemistry Answers
 - Setting Reading Goals Section Tour Of The Periodic Table Holt Chemistry Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Section Tour Of The Periodic Table Holt Chemistry Answers
 - Fact-Checking eBook Content of Section Tour Of The Periodic Table Holt Chemistry Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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