

2.1

CHEMICAL REACTIONS

Section Review

Objectives

1. Measure enthalpy using a calorimeter.
2. Identify and predict chemical reactions using phase diagrams.
3. Apply the law of conservation of energy to chemical reactions.

Vocabulary

- | | | |
|------------------|---------------|----------------------------------|
| 1. endothermic | 2. exothermic | 3. enthalpy |
| 4. heat capacity | 5. joule | 6. law of conservation of energy |

Part A Completion

Use the chapter content to identify a vocabulary word, complete, and write a brief definition. Review the chapter content carefully before doing these problems for practice.

- | | |
|---|-----------|
| 1. Enthalpy change depends on the change in _____. | 1. _____ |
| 2. Enthalpy change is the change in a substance's _____. | 2. _____ |
| 3. Heat capacity is a _____ change. It is measured _____. | 3. _____ |
| 4. Heat _____ is the change in heat of a substance. | 4. _____ |
| 5. _____ is the change in heat of a substance. | 5. _____ |
| 6. _____ is the change in heat of a substance. | 6. _____ |
| 7. _____ is the change in heat of a substance. | 7. _____ |
| 8. _____ is the change in heat of a substance. | 8. _____ |
| 9. _____ is the change in heat of a substance. | 9. _____ |
| 10. _____ is the change in heat of a substance. | 10. _____ |
| 11. _____ is the change in heat of a substance. | 11. _____ |
| 12. _____ is the change in heat of a substance. | 12. _____ |
| 13. _____ is the change in heat of a substance. | 13. _____ |
| 14. _____ is the change in heat of a substance. | 14. _____ |
| 15. _____ is the change in heat of a substance. | 15. _____ |
| 16. _____ is the change in heat of a substance. | 16. _____ |
| 17. _____ is the change in heat of a substance. | 17. _____ |
| 18. _____ is the change in heat of a substance. | 18. _____ |
| 19. _____ is the change in heat of a substance. | 19. _____ |
| 20. _____ is the change in heat of a substance. | 20. _____ |

Part B True-False

Indicate whether each statement is true or false. If false, explain why.

1. The heat capacity of a substance is constant.
2. The heat capacity of a substance is constant.
3. The heat capacity of a substance is constant.
4. The heat capacity of a substance is constant.
5. The heat capacity of a substance is constant.
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17. The heat capacity of a substance is constant.
18. The heat capacity of a substance is constant.
19. The heat capacity of a substance is constant.
20. The heat capacity of a substance is constant.

Section Review 11describing Chemical Reactions

**Linda Swisher, Kevin T. Patton, Gary A.
Thibodeau**



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Table of Contents Section Review 11describing Chemical Reactions

1. Understanding the eBook Section Review 11describing Chemical Reactions
 - The Rise of Digital Reading Section Review 11describing Chemical Reactions
 - Advantages of eBooks Over Traditional Books
2. Identifying Section Review 11describing Chemical Reactions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section Review 11describing Chemical Reactions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section Review 11describing Chemical Reactions
 - Personalized Recommendations
 - Section Review 11describing Chemical Reactions User Reviews and Ratings
 - Section Review 11describing Chemical Reactions and Bestseller Lists

5. Accessing Section Review 11describing Chemical Reactions Free and Paid eBooks
 - Section Review 11describing Chemical Reactions Public Domain eBooks
 - Section Review 11describing Chemical Reactions eBook Subscription Services
 - Section Review 11describing Chemical Reactions Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Section Review 11describing Chemical Reactions Compatibility with Devices
 - Section Review 11describing Chemical Reactions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section Review 11describing Chemical Reactions
 - Highlighting and Note-Taking Section Review 11describing Chemical Reactions
 - Interactive Elements Section Review 11describing Chemical Reactions
8. Staying Engaged with Section Review 11describing Chemical Reactions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section Review 11describing Chemical Reactions
9. Balancing eBooks and Physical Books Section Review 11describing Chemical Reactions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section Review 11describing Chemical Reactions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Section Review 11describing Chemical Reactions
 - Setting Reading Goals Section Review 11describing Chemical Reactions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Section Review 11describing Chemical Reactions
 - Fact-Checking eBook Content of Section Review 11describing Chemical Reactions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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