

Percent Yield Practice

3 pages + 3-page Answer KEY + Periodic Table

Chemistry

Name _____ Date _____

PERCENT YIELD

Show all work and circle or highlight your final answer.

$$\text{percent yield (\%)} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

- 1) 10.0 grams of sodium reacts with excess chlorine gas.
 $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$
The theoretical yield of sodium chloride is 25.0 grams. The actual yield was 23.0 grams. Calculate the percent yield of sodium chloride.

- 2) Hydrogen gas (H_2) reacts with oxygen gas (O_2) to produce water (H_2O). 5.0 grams of hydrogen reacts with excess oxygen gas.
 $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
What is the theoretical yield?

What is the theoretical yield?

- 3) 4.0 grams of carbon reacts with excess oxygen gas.
 $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$
The theoretical yield of carbon dioxide is 25.0 grams. Calculate the percent yield of carbon dioxide.

- 4) The actual yield was 24.0 grams. Calculate the percent yield.

- 5) Magnesium (Mg) reacts with hydrochloric acid (HCl) to produce magnesium chloride (MgCl_2) and hydrogen gas (H_2).
 $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$
What is the theoretical yield?

What is the theoretical yield?

- 6) 1.0 grams of zinc reacts with excess sulfur.
 $\text{Zn} + \text{S} \rightarrow \text{ZnS}$
The theoretical yield of zinc sulfide is 15.0 grams. Calculate the percent yield of zinc sulfide.

- 7) 1.0 grams of phosphorus reacts with excess chlorine.
 $\text{P}_4 + 6\text{Cl}_2 \rightarrow 4\text{PCl}_3$
The theoretical yield of phosphorus trichloride is 4.0 grams. Calculate the percent yield of phosphorus trichloride.

- 8) The actual yield was 40.0 grams. Calculate the percent yield.

- 9) Calcium carbonate (CaCO_3) reacts with hydrochloric acid (HCl) to produce calcium chloride (CaCl_2), carbon dioxide (CO_2), and water (H_2O). 10.0 grams of calcium carbonate reacts with excess hydrochloric acid.
 $\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{CO}_2 + \text{H}_2\text{O}$
What is the theoretical yield of calcium chloride?

$$\begin{array}{r} 10.0 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} = 0.100 \text{ mol CaCO}_3 \\ 0.100 \text{ mol CaCO}_3 \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} = 0.100 \text{ mol CaCl}_2 \\ 0.100 \text{ mol CaCl}_2 \times \frac{147.02 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 14.702 \text{ g CaCl}_2 \end{array}$$

Molar Mass CaCl_2
 $\text{Ca} = 40.08$
 $\text{Cl} = 35.45$
 $\text{Ca} = 40.08$
 $2 \times 35.45 = 70.90$
 $40.08 + 70.90 = 110.98$
 110.98 g

$$\begin{array}{r} 10.0 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCl}_2}{100.09 \text{ g CaCO}_3} = 0.100 \text{ mol CaCl}_2 \\ 0.100 \text{ mol CaCl}_2 \times \frac{147.02 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 14.702 \text{ g CaCl}_2 \end{array}$$

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 $40.08 + 70.90 = 110.98$
 110.98 g

- 10) The actual yield was 15.0 grams. Calculate the percent yield of calcium chloride.

$$\text{percent yield (\%)} = \frac{15.0 \text{ g}}{14.702 \text{ g}} \times 100 = 1.0203 \times 100 = 102.03\% \text{ CaCl}_2$$

- 11) Potassium (K) reacts with iodine (I_2) to produce potassium iodide (KI). 5.0 grams of potassium reacts with excess iodine.
 $2\text{K} + \text{I}_2 \rightarrow 2\text{KI}$
What is the theoretical yield?

$$\begin{array}{r} 5.0 \text{ g K} \times \frac{1 \text{ mol K}}{39.10 \text{ g K}} = 0.128 \text{ mol K} \\ 0.128 \text{ mol K} \times \frac{1 \text{ mol KI}}{2 \text{ mol K}} = 0.064 \text{ mol KI} \\ 0.064 \text{ mol KI} \times \frac{166.10 \text{ g KI}}{1 \text{ mol KI}} = 10.630 \text{ g KI} \end{array}$$

Molar Mass KI
 $\text{K} = 39.10$
 $\text{I} = 126.90$
 $\text{K} = 39.10$
 $\text{I} = 126.90$
 $39.10 + 126.90 = 166.00$
 166.00 g

ANSWER KEY INCLUDED



Percent Yield Practice Problems Answer Key

SJ Ball



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