

Annotated periodic table of the elements

<p>Group names*</p> <ul style="list-style-type: none"> 1 Hydrogen & the alkali metals 2 Alkaline earth metals 3 Lanthanoids 4 Actinoids 5 Boron group 6 Chalcogens 7 Halogens 8 Helium group 9–10 Crystallogens 11–12 Postcogens 13–14 Chalcogens 15–16 Halogens 17 Helium group <p>*Groups 3 to 10 are named after their first members (i.e. Group 3 is the Scandium group)</p>																		
<p>Main group (outer shell electrons, electronegativity, Mendeleev's ideas)</p> <p>Atomic number and symbol, atomic weight and name</p>																		
1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	
H 1.01 Hydrogen	He 4.003 Helium	Li 6.94 Lithium	Be 9.012 Beryllium	B 10.81 Boron	C 12.01 Carbon	N 14.01 Nitrogen	O 16.00 Oxygen	F 18.99 Fluorine	Ne 20.18 Neon	Na 22.99 Sodium	Mg 24.31 Magnesium	Al 26.98 Aluminum	Si 28.09 Silicon	P 30.97 Phosphorus	S 32.06 Sulfur	Cl 35.45 Chlorine	Ar 39.95 Argon	
K 39.10 Potassium	Ca 40.08 Calcium	Sc 44.96 Scandium	Ti 47.87 Titanium	V 50.94 Vanadium	Cr 52.00 Chromium	Mn 54.94 Manganese	Fe 55.85 Iron	Cobalt 58.93 Cobalt	Ni 58.69 Nickel	Cu 63.55 Copper	Zn 65.38 Zinc	Ga 69.72 Gallium	Ge 72.61 Germanium	As 74.92 Arsenic	Se 78.96 Selenium	Br 79.90 Bromine	Kr 83.80 Krypton	
Rb 85.47 Rubidium	Sr 87.62 Strontium	Y 88.91 Yttrium	Zr 91.22 Zirconium	Nb 92.91 Niobium	Mo 95.94 Molybdenum	Tc 98.91 Technetium	Ru 101.07 Ruthenium	Rh 101.07 Rhodium	Pd 106.42 Palladium	Ag 107.87 Silver	Cd 112.41 Cadmium	In 114.82 Indium	Tl 204.38 Thallium	Pb 208.28 Lead	Hg 200.59 Mercury	Tl 204.38 Thallium	Pb 208.28 Lead	
Cs 132.91 Cesium	Ba 137.33 Barium	La 138.91 Lanthanum	Hf 178.49 Hafnium	Ta 180.95 Tantalum	W 183.84 Tungsten	Re 186.21 Rhenium	Os 190.23 Osmium	Ir 192.22 Iridium	Pt 195.08 Platinum	Au 196.97 Gold	Hg 200.59 Mercury	Tl 204.38 Thallium	Pb 208.28 Lead	Bi 208.98 Bismuth	Po 209 Polonium	At 210 Astatine	Rn 222 Radon	
Fr 223 Francium	Ra 226 Radium	57 to 70 Lanthanoids	Lr 260 Lawrencium	Rf 261 Rutherfordium	Hf 261 Hafnium	Sg 263 Seaborgium	Bh 264 Bohrium	Hs 265 Hassium	Mt 266 Meitnerium	Ds 268 Darmstadtium	Rg 269 Roentgenium	Cn 284 Copernicium	Nh 285 Nihonium	Fl 286 Flerovium	Mc 287 Moscovium	Lv 289 Livermorium	Ts 289 Tennessine	Og 284 Oganesson
89 to 102 Actinoids																		

CLASS COUNTERPARTS	
Metals	Nonmetals
Active	Inert
Transition	Intermetallic
Fracture	Metallic
Node (Ru-Pt, Os-Au)	Node gas

La 138.91 Lanthanum	Ce 140.12 Cerium	Pr 140.91 Praseodymium	Nd 144.24 Neodymium	Pm 144.91 Promethium	Sm 150.36 Samarium	Eu 151.96 Europium	Gd 157.25 Gadolinium	Tb 158.93 Terbium	Dy 162.50 Dysprosium	Ho 164.93 Holmium	Er 167.26 Erbium	Tm 168.93 Thulium	Yb 173.05 Ytterbium
Ac 227.03 Actinium	Th 232.04 Thorium	Pa 231.04 Protactinium	U 238.03 Uranium	Np 237.05 Neptunium	Pu 244.06 Plutonium	Am 243.06 Americium	Cm 247.07 Curium	Bk 247.07 Berkelium	Cf 251.08 Californium	Es 252.08 Einsteinium	Fm 257.10 Fermium	Md 258.10 Mendelevium	No 259 Nobelium

Radioactive

Openness is likely to be either a metalloid or a post-transition metal

Active metals

Pre-transition metals
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).

Rare earths
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale-colored compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.

Actinides
Soft, dense and reactive metals. Those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form colored compounds.

Transition metals
Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

Noble metals
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very weakly electropositive; inertness oxides.

Frontier metals
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterized, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and bismuthates.

Metalloids
Metallic looking brittle, fair electrical conductors. All known in amorphous forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals
H and N are colorless gases, C, P, and Se are metalloids looking solids while Si is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic or near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).

Group 1
The eight in Si-P, and two otherwise (in a La branch - La-Cu-Ac, and 4 Lu branch - La-Lu). Trends going from Si-P, and the La branch resemble those of groups 1-2.

Coactive nonmetals
O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically coactive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

Noble gases
Colorless, odorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

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David R. Klein



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Table of Contents Periodic Table Guide With Answer

1. Understanding the eBook Periodic Table Guide With Answer
 - The Rise of Digital Reading Periodic Table Guide With Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Guide With Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Guide With Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Guide With Answer
 - Personalized Recommendations
 - Periodic Table Guide With Answer User Reviews and Ratings
 - Periodic Table Guide With Answer and Bestseller Lists

5. Accessing Periodic Table Guide With Answer Free and Paid eBooks
 - Periodic Table Guide With Answer Public Domain eBooks
 - Periodic Table Guide With Answer eBook Subscription Services
 - Periodic Table Guide With Answer Budget-Friendly Options
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 - Periodic Table Guide With Answer Compatibility with Devices
 - Periodic Table Guide With Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table Guide With Answer
 - Highlighting and Note-Taking Periodic Table Guide With Answer
 - Interactive Elements Periodic Table Guide With Answer
8. Staying Engaged with Periodic Table Guide With Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Guide With Answer
9. Balancing eBooks and Physical Books Periodic Table Guide With Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Guide With Answer
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table Guide With Answer
 - Setting Reading Goals Periodic Table Guide With Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table Guide With Answer
 - Fact-Checking eBook Content of Periodic Table Guide With Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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