

## Chapter 4 and 5 Study Guide (Atomic Structure and the Periodic Table)

1. What did Dalton's model look like? List the parts of his atomic theory.

His model of an atom was a solid sphere.



Parts of the atomic theory:

- All matter is made up of individual particles called atoms, which can't be divided.
  - All elements made of atoms
  - Atoms of same element have same mass, atoms of different elements have different masses
  - Compounds have more than one type of element
  - Atoms of different elements always combine the same way in compound

2. What did JJ Thompson discover? Draw/describe his model.

JJ Thompson discovered the electron. "Plum pudding model" evenly scattered charges.

3. What did Millikan discover and draw/describe his model.

Millikan discovered the exact charge and mass of an electron. His model was of an atom was that there were negative charges scattered in a positive substance.

4. What did Rutherford discover? Draw/describe his model.

Discovered the dense, positive center of an atom. Model was that protons were in center, electrons moved on the outside.

5. What did Bohr discover? Draw/describe his model.

Electrons moved in energy levels. Nucleus at center, electrons moving in energy levels on the outside.

6. To find the atomic mass what two subatomic particles do you add together?

Protons + neutrons

7. Atomic number equals the number of \_\_\_\_\_ protons \_\_\_\_\_ and \_\_\_\_\_ electrons \_\_\_\_\_.

8. What are the three main parts to the atom? Name them, their charge, location and their mass.

Particle	Charge	Location	Mass
Electron	-	Outside nucleus	0 amu
Proton	+	Nucleus	1 amu
Neutron	0	Nucleus	1 amu

# Periodic Table Test With Answer Key

**Amar Bhutani**



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## **Table of Contents Periodic Table Test With Answer Key**

1. Understanding the eBook Periodic Table Test With Answer Key
  - The Rise of Digital Reading Periodic Table Test With Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Test With Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Table Test With Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Test With Answer Key
  - Personalized Recommendations
  - Periodic Table Test With Answer Key User Reviews and Ratings
  - Periodic Table Test With Answer Key and Bestseller Lists
5. Accessing Periodic Table Test With Answer Key Free and Paid eBooks
  - Periodic Table Test With Answer Key Public Domain eBooks
  - Periodic Table Test With Answer Key eBook Subscription Services
  - Periodic Table Test With Answer Key Budget-Friendly Options

6. Navigating Periodic Table Test With Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Periodic Table Test With Answer Key Compatibility with Devices
  - Periodic Table Test With Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Table Test With Answer Key
  - Highlighting and Note-Taking Periodic Table Test With Answer Key
  - Interactive Elements Periodic Table Test With Answer Key
8. Staying Engaged with Periodic Table Test With Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Table Test With Answer Key
9. Balancing eBooks and Physical Books Periodic Table Test With Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Table Test With Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table Test With Answer Key
  - Setting Reading Goals Periodic Table Test With Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table Test With Answer Key
  - Fact-Checking eBook Content of Periodic Table Test With Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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