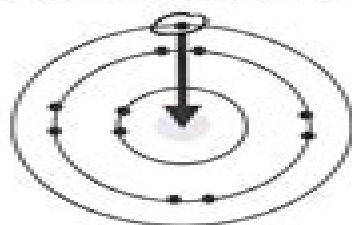
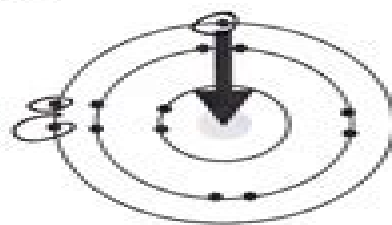
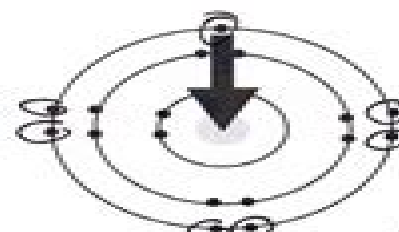


Model 4 – Period 3 Elements

Sodium $11p^+$ Aluminum $13p^+$ Chlorine $17p^+$

13. Using the periodic table, locate the elements whose atoms are diagrammed in Model 4. Are the elements in the same column or the same row?

The same (3rd) row. ← Horizontal!

14. Circle the outermost electron(s) in each of the atoms in Model 4.

15. Which of the three atoms diagrammed in Model 4 has the strongest attraction for its outermost electron(s)? *Chlorine*



16. Consider the information in Model 4.

- a. As you move from the smallest atom to the largest atom, does the distance between the outermost electron(s) and the nucleus change significantly?

The distance changes very little if at all.

- b. Can the differences in the attractive force shown by the arrows be explained by a change in the distance between the electron(s) and the nucleus?

If the distance does not change dramatically, the change in force cannot be attributed to a change in distance.

- c. On the diagrams in Model 4, write the number of protons located in the nucleus of each atom.

- d. Can the differences in attractive forces shown by the arrows in Model 4 be explained by a change in the number of protons in the nucleus? If yes, explain the relationship in Model 4.

Yes, the force increases with an increase of protons.



17. For each set of elements below, circle the element whose atoms will have a stronger attractive force between their outermost electron(s) and the nucleus.

a. Ba and Cs

b. Cr and Cu

c. Al and Xe

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VM Jensen



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