

Graphing Periodic Trends Activity

Name: _____ Period: _____

Use the information in these tables to complete the graph as described below

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
H	31	1312	2.1
Li	128	520	1.0
Na	166	496	0.9
K	203	410	0.8
Rb	220	403	0.8
Cs	244	376	0.7

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
Na	166	496	0.9
Mg	141	738	1.2
Al	121	578	1.5
Si	111	787	1.8
P	107	1012	2.1
S	105	1000	2.5
Cl	102	1251	3.0
Ar	106	1521	---

1. Using colored pencils or pen, list under the symbol (in this order!) the Atomic Radius, First Ionization Energy, and Electronegativity.
2. Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
4. Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)

H	<p>Across a period (→) atomic radius tends to _____.</p> <p>Across a period (→) first ionization energy tends to _____.</p> <p>Across a period (→) electronegativity tends to _____.</p>						
Li							
Na 166 496 0.9	Mg	Al	Si	P	S	Cl	Ar
K	<p>Down a Group (↓) atomic radius tends to _____.</p> <p>Down a Group (↓) first ionization energy tends to _____.</p> <p>Down a Group (↓) electronegativity tends to _____.</p>						
Rb							
Cs							

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Table of Contents Periodic Trends Practice Answer Key

1. Understanding the eBook Periodic Trends Practice Answer Key
 - The Rise of Digital Reading Periodic Trends Practice Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Practice Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Practice Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Practice Answer Key
 - Personalized Recommendations

- Periodic Trends Practice Answer Key User Reviews and Ratings
- Periodic Trends Practice Answer Key and Bestseller Lists
- 5. Accessing Periodic Trends Practice Answer Key Free and Paid eBooks
 - Periodic Trends Practice Answer Key Public Domain eBooks
 - Periodic Trends Practice Answer Key eBook Subscription Services
 - Periodic Trends Practice Answer Key Budget-Friendly Options
- 6. Navigating Periodic Trends Practice Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Practice Answer Key Compatibility with Devices
 - Periodic Trends Practice Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Practice Answer Key
 - Highlighting and Note-Taking Periodic Trends Practice Answer Key
 - Interactive Elements Periodic Trends Practice Answer Key
- 8. Staying Engaged with Periodic Trends Practice Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Practice Answer Key
- 9. Balancing eBooks and Physical Books Periodic Trends Practice Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Practice Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Practice Answer Key
 - Setting Reading Goals Periodic Trends Practice Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Practice Answer Key
 - Fact-Checking eBook Content of Periodic Trends Practice Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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