



COVALENT BONDING

Practice Problems

Do your answers to these very practice problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following compounds as ionic or covalent.

- a. Na_2O (ionic) b. CH_4 (covalent) c. HCl (covalent)
 d. KNO_3 (ionic) e. H_2O (covalent)

2. Write the formula for the following molecular compounds.

- a. carbon dioxide b. sulfur dioxide c. carbon monoxide
 d. N_2O e. SO_2 f. CO

3. Write the formula for the following molecular compounds: nitrogen.

4. Write the formula for the following molecular compounds: carbon dioxide, carbon monoxide, carbon disulfide, carbon tetrachloride, sulfur dioxide, sulfur trioxide, sulfur hexafluoride, and phosphorus pentachloride.

5. Write the formula for the following molecular compounds: dinitrogen, carbon monoxide, carbon dioxide, carbon disulfide, carbon tetrachloride, sulfur dioxide, sulfur trioxide, sulfur hexafluoride, and phosphorus pentachloride.

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for hydrogen fluoride, HF .



2. Draw the Lewis structure for methane, CH_4 .



3. Draw the Lewis structure for carbon dioxide, CO_2 .



4. Draw the Lewis structure for acetylene, C_2H_2 .

5. Draw the Lewis structure for the following: dinitrogen, carbon monoxide, carbon dioxide, carbon disulfide, carbon tetrachloride, sulfur dioxide, sulfur trioxide, sulfur hexafluoride, and phosphorus pentachloride.



6. Draw the Lewis structure for the following: dinitrogen, carbon monoxide, carbon dioxide, carbon disulfide, carbon tetrachloride, sulfur dioxide, sulfur trioxide, sulfur hexafluoride, and phosphorus pentachloride.

SECTION 8.3 BONDING: THE ORBITAL

1. Write the formula for the bond between the carbon atoms in acetylene, C_2H_2 . The bond between the two carbon atoms is a triple bond.

2. Write the formula for the bond between the two carbon atoms in ethane, C_2H_6 . The bond between the two carbon atoms is a single bond.

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