

Chem 1020 Molarity worksheet, version 2

1. What is the molarity of a solution that contains 5.83 g LiCl in 1.50 L solution?

$$5.83 \text{ g LiCl} \times \frac{1 \text{ mol LiCl}}{42.389 \text{ g LiCl}} = 0.13754 \text{ mol LiCl}$$

$$M = \frac{\text{mol solute}}{\text{L solution}} = \frac{0.13754 \text{ mol LiCl}}{1.50 \text{ L solution}} = 0.0917 \text{ M}$$

2. What is the volume of a 0.300 M CuSO<sub>4</sub> solution that contains 12.4 g CuSO<sub>4</sub>?

$$12.4 \text{ g CuSO}_4 \times \frac{1 \text{ mol CuSO}_4}{159.6 \text{ g CuSO}_4} \times \frac{1 \text{ L solution}}{0.300 \text{ mol CuSO}_4} = 0.26 \text{ L solution}$$

3. What mass of solid KBr would be needed to prepare 750.0 mL of 1.75 M KBr solution? How specifically would this solution be prepared?

$$750.0 \text{ mL KBr soln} \times \frac{1 \text{ L soln}}{1000 \text{ mL soln}} \times \frac{1.75 \text{ mol KBr}}{1 \text{ L soln}} \times \frac{119.01 \text{ g KBr}}{1 \text{ mol soln}} = 156 \text{ g KBr}$$

Prepare this solution by measuring 156 g KBr into a 750 mL volumetric flask, and adding enough water to make 750 mL solution.

4. What is the % by mass of 28.3 g of solution which contains 3.56 g NaCl?

$$\% (\text{mass}) = \frac{\text{mass solute}}{\text{mass solution}} \times 100\% = \frac{3.56 \text{ g NaCl}}{28.3 \text{ g solution}} \times 100\% = 12.6\%$$

5. What volume of methanol is required to prepare 2.0 L of a 15.0% by volume solution?

$$\% (\text{vol}) = \frac{\text{vol solute}}{\text{vol solution}} \times 100\%$$

$$\text{vol solute} = \left( \frac{\% (\text{vol})}{100\%} \right) \times \text{vol solution} = \left( \frac{15.0\%}{100\%} \right) \times 2.0 \text{ L} = 0.30 \text{ L}$$

6. What volume of 12.0 M HCl is required to make 500.0 mL of 0.400 M HCl? How specifically would this solution be prepared?

$$M_1 V_1 = M_2 V_2$$

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(0.400 \text{ M})(500.0 \text{ mL})}{(12.0 \text{ M})} = 16.7 \text{ mL}$$

Prepare this solution by measuring 16.7 mL of 12.0 M HCl into a 500 mL volumetric flask, and adding water to make 500 mL of solution.

7. If 17.0 mL of 16.0 M H<sub>2</sub>SO<sub>4</sub> solution is diluted to 100.0 mL, what is the diluted concentration?

$$M_1 V_1 = M_2 V_2$$

$$M_2 = \frac{M_1 V_1}{V_2} = \frac{(16.0 \text{ M})(17.0 \text{ mL})}{(100.0 \text{ mL})} = 2.72 \text{ M}$$

# Section Concentration And Molarity Quiz Answer Key

**Eugene Meyer**



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## **Table of Contents Section Concentration And Molarity Quiz Answer Key**

1. Understanding the eBook Section Concentration And Molarity Quiz Answer Key
  - The Rise of Digital Reading Section Concentration And Molarity Quiz Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Section Concentration And Molarity Quiz Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Section Concentration And Molarity Quiz Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Section Concentration And Molarity Quiz Answer Key

- Personalized Recommendations
  - Section Concentration And Molarity Quiz Answer Key User Reviews and Ratings
  - Section Concentration And Molarity Quiz Answer Key and Bestseller Lists
5. Accessing Section Concentration And Molarity Quiz Answer Key Free and Paid eBooks
    - Section Concentration And Molarity Quiz Answer Key Public Domain eBooks
    - Section Concentration And Molarity Quiz Answer Key eBook Subscription Services
    - Section Concentration And Molarity Quiz Answer Key Budget-Friendly Options
  6. Navigating Section Concentration And Molarity Quiz Answer Key eBook Formats
    - ePub, PDF, MOBI, and More
    - Section Concentration And Molarity Quiz Answer Key Compatibility with Devices
    - Section Concentration And Molarity Quiz Answer Key Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Section Concentration And Molarity Quiz Answer Key
    - Highlighting and Note-Taking Section Concentration And Molarity Quiz Answer Key
    - Interactive Elements Section Concentration And Molarity Quiz Answer Key
  8. Staying Engaged with Section Concentration And Molarity Quiz Answer Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Section Concentration And Molarity Quiz Answer Key
  9. Balancing eBooks and Physical Books Section Concentration And Molarity Quiz Answer Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Section Concentration And Molarity Quiz Answer Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Section Concentration And Molarity Quiz Answer Key
    - Setting Reading Goals Section Concentration And Molarity Quiz Answer Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Section Concentration And Molarity Quiz Answer Key

- Fact-Checking eBook Content of Section Concentration And Molarity Quiz Answer Key
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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