



Chemical Changes

Directions: Read the scenarios below and answer the following questions.



1. What are the products of this chemical reaction?

the products are the substances that are formed at the end of the chemical reaction, the
products are silver and copper(II) chloride. 2Ag and CuCl₂

2. What is the reactant in this chemical reaction? Is it a solid, liquid, gas, or aqueous solution? What is the physical state of this reactant? It is a solid metal.

3. What is the physical state of the reactant? Is it a solid, liquid, gas, or aqueous solution?

the reactant is copper metal. It is a solid metal. The products are silver metal and copper(II) chloride. The products are a solid metal and an aqueous solution. The products are silver metal and copper(II) chloride. The products are silver metal and copper(II) chloride. The products are silver metal and copper(II) chloride.

4. According to the law of conservation of mass, if 10.0 grams of silver metal and 5.0 grams of copper metal react, how much silver metal and copper(II) chloride will be produced?

15 grams



5. What is the reactant in this chemical reaction? Is it a solid, liquid, gas, or aqueous solution?

6. What is the physical state of the reactant? Is it a solid, liquid, gas, or aqueous solution?

7. What are the physical states of the products? Is it a solid, liquid, gas, or aqueous solution?

the products are liquid water and gas hydrogen.

8. What is the physical state of the product? Is it a solid, liquid, gas, or aqueous solution?

9. What is the physical state of the product? Is it a solid, liquid, gas, or aqueous solution?

10. According to the law of conservation of mass, if 2.0 grams of hydrogen gas and 16.0 grams of oxygen gas react, how much water will be produced?

18.0 grams

Section Reinforcement Chemical Changes Answers

Rosina Ehmman



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