

## PERIODIC TABLE TRENDS

**Step 1:** Complete the squares for each element by adding the atomic number, name, and atomic mass.

Write the atomic number at the top of the square.

Write the atomic name under the symbol.

Write the atomic mass at the bottom of the square.

**Step 2:** Determine the number of electrons, protons, and neutrons in each statement.

**Step 3:** Create a Bohr diagram for each element.

**Step 4:** Draw a Lewis Dot Structure for each element.

**Step 5:** Show whether the element is a metal, nonmetal, or metalloid by circling the correct response. Also determine if the element is a solid or gas by circling the correct choice.

**Step 6:** Use the following colors to shade in the square for each element. You should **ONLY** color in the small square in the upper left-hand corner and not the entire card.

Green – Li and Na

Purple – F and Cl

Brown – N and P

Pink – O and S

Orange – B and Al

Yellow – He, Ne, and Ar

Blue – Be and Mg

Red – C and Si

White – H

**Step 7:** Cut the cards apart and arrange according to atomic number in the pattern shown below. Once you have the cards arranged in the correct order, glue them to a large sheet of construction paper.

1							2
3	4	5	6	7	8	9	10
11	12	13	14	15	16	17	18

**Step 8:** Put a title on your table.

**Step 9:** Make a key for each color with the name of the group for that column.

**Step 10:** Write the column number at the top of each group and the period number on the left side of the table. Be careful, column numbers are not in order from 1-18 and the transition metals, columns 3-12 are not on the table.

**Step 11:** Draw the red zigzag line that separates the metals from the nonmetals.

**Step 12:** Answer the questions on the worksheet using the information on your periodic table.

\_\_\_\_\_

Be

\_\_\_\_\_

P = \_\_\_\_\_

N = \_\_\_\_\_

E = \_\_\_\_\_



Bohr Diagram



Dot Structure

Metal

Solid

Nonmetal

Gas

Metalloid

Gas

# Section Trends In The Periodic Table Answer Key

**Kaplan Test Prep**



## Section Trends In The Periodic Table Answer Key:

**Kaplan SAT Subject Test Chemistry 2015-2016** Kaplan Test Prep, 2015-03-03 Essential strategies practice and review to ace the SAT Subject Test Chemistry Getting into a top college has never been more difficult Students need to distinguish themselves from the crowd and scoring well on a SAT Subject Test gives students a competitive edge Kaplan's SAT Subject Test Chemistry is the most up to date guide on the market with complete coverage of both the content review and strategies students need for success on test day Kaplan's SAT Subject Test Chemistry features A full length diagnostic test Full length practice tests Focused chapter summaries highlights and quizzes Detailed answer explanations Proven score raising strategies End of chapter quizzes Kaplan is serious about raising students scores we guarantee students will get a higher score

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***Chemistry for the IB Diploma Workbook with CD-ROM*** Jacqueline Paris, 2017-04-06 Chemistry for the IB Diploma Second edition covers in full the requirements of the IB syllabus for Chemistry for first examination in 2016 This workbook is specifically for the IB Chemistry syllabus for examination from 2016 The Chemistry for the IB Diploma Workbook contains straightforward chapters that build learning in a gradual way first outlining key terms and then providing students with plenty of practice questions to apply their knowledge Each chapter concludes with exam style questions This structured approach reinforces learning and actively builds students confidence using key scientific skills handling data evaluating information and problem solving This helps empower students to become confident and independent learners Answers to all of the questions are on the CD ROM

**Ebook: Introductory Chemistry: An Atoms First Approach** Burdge, 2016-04-16 Ebook Introductory Chemistry An Atoms First Approach

**Metal Complexes in Aqueous Solutions** Arthur E. Martell, Robert D.

Hancock,2013-06-29 Stability constants are fundamental to understanding the behavior of metal ions in aqueous solution. Such understanding is important in a wide variety of areas such as metal ions in biology, biomedical applications, metal ions in the environment, extraction, metallurgy, food chemistry, and metal ions in many industrial processes. In spite of this importance, it appears that many inorganic chemists have lost an appreciation for the importance of stability constants and the thermodynamic aspects of complex formation with attention focused over the last thirty years on newer areas such as organometallic chemistry. This book is an attempt to show the richness of chemistry that can be revealed by stability constants when measured as part of an overall strategy aimed at understanding the complexing properties of a particular ligand or metal ion. Thus, for example, there are numerous crystal structures of the Li ion with crown ethers. What do these indicate to us about the chemistry of Li with crown ethers? In fact, most of these crystal structures are in a sense misleading in that the Li ion forms no complexes or at best very weak complexes with familiar crown ethers such as 12-crown-4 in any known solvent. Thus, without the stability constants, our understanding of the chemistry of a metal ion with any particular ligand must be regarded as incomplete. In this book, we attempt to show how stability constants can reveal factors in ligand design which could not readily be deduced from any other physical technique.

**A-Level Chemistry for AQA: Year 1 & 2 Student Book**, 2020-10-05 This fantastic CGP Student Book comprehensively covers both years of AQA A Level Chemistry. It is bursting with in-depth, accessible notes explaining every course topic plus all of the Required Practicals. Everything is supported by clear diagrams, photographs, tips, and worked examples. Throughout the book, there are lots of practice questions and exam-style questions with answers at the back. There is detailed guidance on Maths Skills and Practical Skills as well as indispensable advice for success in the final exams. If you'd prefer Year 1, 9781782943211, Year 2, 9781782943266, in separate books, CGP has them too. And for more detailed coverage of the mathematical elements of A Level Chemistry, try our Essential Maths Skills book, 978182944720.

**Ions in Solution** J Burgess, 1999-10-01 This outline of the principles and chemical interactions in inorganic solution chemistry delivers a course module in an area of considerable complexity. Problems with solutions and tutorial hints to test comprehension have been added as a feature to check readers' understanding and assist self-study. Exercises and projects are also provided to help readers deepen and extend their knowledge and understanding. Inorganic solution chemistry is treated thoroughly. Emphasis is placed upon NMR, UV-VIS, IR, Raman spectroscopy, X-ray diffraction, and such topics as acid-base behaviour, stability constants, and kinetics.

**Holt Chemistry** Holt, Rinehart & Winston, 2003-01-24 *9PPChe2010to2017 Urdu Tube*, 9PPChe2010to2017 [Hands-On General Science Activities With Real-Life Applications](#) Pam Walker, Elaine Wood, 2008-04-21 In this second edition of *Hands-On General Science Activities with Real-Life Applications*, Pam Walker and Elaine Wood have completely revised and updated their must-have resource for science teachers of grades 5-12. The book offers a dynamic collection of classroom-ready lessons, projects, and lab activities that encourage students to integrate basic science concepts and skills into everyday life.

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**Solutions Manual to Accompany Inorganic Chemistry** Alen Hadzovic, 2018 As you master each chapter in Inorganic Chemistry having detailed solutions handy allows you to confirm your answers and develop your ability to think through the problem solving process

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