

## Stoichiometry Reflect

Learning Target... I can use stoichiometry to calculate amounts

1a. Balance this equation

1b. How many moles of potassium nitride was used if the student weighed out 15.6 g of  $\text{K}_3\text{N}$ ?

$$15.6 \text{ g } \text{K}_3\text{N} \left( \frac{1 \text{ mole}}{131.31 \text{ g}} \right) = \underline{0.119 \text{ moles } \text{K}_3\text{N}}$$

1c. What is the amount of moles of potassium oxide produced in this reaction?

$$0.119 \text{ moles } \text{K}_3\text{N} \left( \frac{3 \overset{\text{moles}}{\text{K}_2\text{O}}}{2 \overset{\text{moles}}{\text{K}_3\text{N}}} \right) = \underline{0.178 \text{ moles } \text{K}_2\text{O}}$$

1d. How many grams of potassium oxide was produced in this reaction?

$$0.178 \text{ moles } \text{K}_2\text{O} \left( \frac{94.20 \text{ g}}{1 \text{ mole}} \right) = \underline{16.8 \text{ g } \text{K}_2\text{O}}$$

2. Using the following chemical reaction:

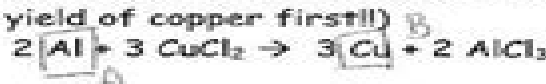


and, knowing that 61.6 g of copper (II) nitrate were used in this experiment, how many grams of mercury were also used?

$$61.6 \text{ g } \text{Cu}(\text{NO}_3)_2 \left( \frac{1 \text{ mole } A}{187.57 \text{ g}} \right) \left( \frac{2 \text{ mol } B}{1 \text{ mol } A} \right) \left( \frac{200.59 \text{ g}}{1 \text{ mol } B} \right) = \underline{132 \text{ g } \text{Hg}}$$

Challenge...

3. What is the percent yield of the following reaction if 2.38 g of aluminum was used in the lab, and 6.15 g of copper was measured at the end of the experiment? (hint: you must calculate the theoretical yield of copper first!)



$$\% \text{ yield} = \frac{6.15 \text{ g}}{8.41 \text{ g}} = \underline{13.1\%}$$

$$2.38 \text{ g } \text{Al} \left( \frac{1 \text{ mol}}{26.98 \text{ g}} \right) \left( \frac{3 \text{ mol } B}{2 \text{ mol } A} \right) \left( \frac{63.55 \text{ g}}{1 \text{ mol}} \right) \Rightarrow 8.41 \text{ g } \text{Cu} = \text{theoretical}$$

Show work on back

# Section Quiz Introduction To Stoichiometry Answers

**Regina M. Murphy**



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