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Stability in Bonding

Directions: Each statement below contains a pair of terms or phrases in parentheses. Circle the term or phrase that makes each statement true.

- The properties of a compound are (the same as, different from) the properties of the elements that make up the compound.
- 2. Na and Cl are chemical (symbols, formulas).
- 3. NaCl and NaOH are chemical (symbols, formulas).
- In the formula H₂O, the number 2 is a (subscript, superscript).
- The number 2 in the formula H₂O tells you that each unit of this compound contains two (hydrogen, oxygen) atoms.
- If a symbol in a chemical formula does not have a subscript after it, a unit of that compound contains (no atoms, one atom) of that element.
- 7. The total number of atoms in Fe₂O₃ is (two, five, six).
- 8. There are (three, seven, ten) different elements in H,SO,.
- 9. An atom is chemically stable if its outer energy level (is filled with, contains no) electrons.
- 10. For atoms of most noble gases and most other elements, the outer energy level is full when it has (3, 8) electrons.
- 11. The noble gases do not readily form compounds because they (are, are not) chemically stable.
- 12. A chemical bond is a (force, chemical) that holds atoms together in a compound.
- 13. Chemical bonds form when atoms lose, gain, or (share, multiply) electrons.

Directions: Complete the table below by using the formula of each compound to identify the elements that each compound contains and the number of atoms of each of these elements in a unit of the compound. The first formula has been done for you.

Formula	Element 1	Element 2	Element 3
H,O	2 hydrogen	1 oxygen	
14. NaOH			
15. NaCl		re district	
16. NH,			
17. H,SO.			
18. SiO,			The second of the second

Section Stability In Bonding Answer Key

Disha Experts

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