

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

- 1. Represent the processes of a reaction.
- 2. Explain how to write balanced equations.
- 3. List the steps for writing complete ionic equations.

Vocabulary

- 1. chemical equation
- 2. balanced equation
- 3. catalyst
- 4. coefficient
- 5. skeletal equation

Part A Completion

Write the coefficient for each in the balanced equation, and balance the chemical equation by changing the coefficients. Write balanced skeletal, complete ionic, and net ionic equations.

1. A chemical equation is always represented by a balanced _____ 1. _____
 The chemical formula for a chemical compound is _____ 2. _____ The mass
 number for an element is represented by the _____ 3. _____ An element's atomic
 mass is represented by _____ 4. _____ A chemical equation is written with
 reactants on the left and products on the right _____ 5. _____ A chemical equation is
 balanced when the number of atoms of each element is the same on both sides of the
 equation. The equation is balanced because the number of _____ 6. _____ on
 the reactant side is the same as the number of _____ 7. _____ on
 the product side.

2. A chemical equation is written with the reactants on the left and the products on the
 right. The equation is balanced when the number of atoms of each element is the same
 on both sides. A chemical equation is balanced when the number of _____ 8. _____ on
 the reactant side is the same as the number of _____ 9. _____ on the
 product side.

- 1. equation
- 2. formula
- 3. mass number
- 4. atomic mass
- 5. balanced
- 6. reactants
- 7. products
- 8. atoms
- 9. molecules

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EW Minium

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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