

Electron Configurations - KEY

Note: The electron configurations in this worksheet assume that lanthanum (La) is the first element in the 4f block and that actinium (Ac) is the first element in the 5f block. If your periodic table doesn't agree with this, your answers for elements near the f-orbitals may be slightly different (one more/less electron for the f-orbital).

- 1) sodium $1s^2 2s^2 2p^6 3s^1$
- 2) iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^5$
 also might be written (due to the close energy of d and f and similarities of elements versus predictable memorized config: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 6d^1 5f^4$
- 6) cobalt $[Ar] 4s^2 3d^7$
- 7) silver $[Kr] 5s^2 4d^9$
- 8) tellurium $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium $[Rn] 7s^2$
- 10) lawrencium $[Rn] 7s^2 5f^{14} 6d^1$
 also might be written (due to the close energy of d and f and similarities of elements versus predictable memorized config: $[Rn] 7s^2 6d^1 5f^{14}$
- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ sulfur
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ rubidium
- 13) $[Kr] 5s^2 4d^{10} 5p^3$ antimony
- 14) $[Xe] 6s^2 4f^{14} 5d^6$ osmium
- 15) $[Rn] 7s^2 5f^{11}$ einsteinium *note: no d's, so the p. table referenced has Ac as 5f*
- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ not valid (take a look at "4d"; must be 3 d not 4!)
- 17) $1s^2 2s^2 2p^6 3s^3 3d^5$ not valid (3p comes after 3s); also no 3 electrons in s
- 18) $[Ra] 7s^2 5f^8$ not valid (radium isn't a noble gas)
- 19) $[Kr] 5s^2 4d^{10} 5p^3$ valid
- 20) $[Xe]$ not valid (an element can't be its own electron configuration)

Section 5 3 Electron Configuration Answers

Jicheng Xie



Section 5 3 Electron Configuration Answers:

Basic Chemistry Concepts and Exercises John Kenkel, 2011-07-08 Chemistry can be a daunting subject for the uninitiated and all too often introductory textbooks do little to make students feel at ease with the complex subject matter *Basic Chemistry Concepts and Exercises* brings the wisdom of John Kenkel's more than 35 years of teaching experience to communicate the fundamentals of chemistry in a practical down to earth manner Using conversational language and logically assembled graphics the book concisely introduces each topic without overwhelming students with unnecessary detail Example problems and end of chapter questions emphasize repetition of concepts preparing students to become adept at the basics before they progress to an advanced general chemistry course Enhanced with visualization techniques such as the first chapter's mythical microscope the book clarifies challenging abstract ideas and stimulates curiosity into what can otherwise be an overwhelming topic Topics discussed in this reader friendly text include Properties and structure of matter Atoms molecules and compounds The Periodic Table Atomic weight formula weights and moles Gases and solutions Chemical equilibrium Acids bases and pH Organic chemicals The appendix contains answers to the homework exercises so students can check their work and receive instant feedback as to whether they have adequately grasped the concepts before moving on to the next section Designed to help students embrace chemistry not with trepidation but with confidence this solid preparatory text forms a firm foundation for more advanced chemistry training

Understanding General Chemistry Atef Korchef, 2022-03-07 *Understanding General Chemistry* details the fundamentals of general chemistry through a wide range of topics relating the structure of atoms and molecules to the properties of matter Written in an easy to understand format with helpful pedagogy to fuel learning the book features main objectives at the beginning of each chapter get smart sections and check your reading section at the end of each chapter The text is filled with examples and practices that illustrate the concepts at hand In addition a summary and extensive MCQs exercises and problems with the corresponding answers and explanations are readily available Additional features include Alerts students to common mistakes and explains in simple ways and clear applications how to avoid these mistakes Offers answers and comments alongside sample problems enabling students to self evaluate their skill level Includes powerful methods easy steps simple and accurate interpretations and engaging applications to help students understand complex principles Provides a bridge to more complex topics such as solid state chemistry organometallic chemistry chemistry of main group elements inorganic chemistry and physical chemistry This introductory textbook is ideal for chemistry courses for non science majors as well as health sciences and preparatory engineering students

Basic Concepts of Chemistry Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader

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Chemistry John Olmsted, Greg Williams, Robert C. Burk, 2020 Chemistry 4th Edition is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers and distinguish this text from other offerings It more accurately reflects the curriculum of most Canadian institutions Chemistry is sufficiently rigorous while engaging and retaining student interest through its accessible language and clear problem solving program without an excess of material and redundancy

Electron Devices and Circuits Atul. P. Godse, Uday A. Bakshi, 2020-11-01 The book covers all the aspects of theory analysis and design of Electron Devices and Circuits for the undergraduate course The concepts of p n junction devices BJT JFET MOSFET electronic devices including UJT thyristors IGBT Amplifier circuits BJT JFET and MOSFET amplifiers multistage and differential amplifiers feedback amplifiers and oscillators are explained comprehensively The book explains various p n junction devices including diode LED laser diode Zener diode and Zener diode regulator The different types of rectifiers are explained in support The book covers the construction operation and characteristics of BJT JFET MOSFET UJT Thyristors SCR Diac and Triac and IGBT It explains the biasing of BJT JFET and MOSFET amplifiers basic BJT JFET and MOSFET amplifiers with h parameters and r parameters equivalent circuits multistage amplifiers differential amplifiers BiCMOS amplifier single tuned amplifiers neutralization methods power amplifiers and frequency response Finally the book incorporates a detailed discussion of the analysis of the current series voltage series current shunt and voltage shunt feedback amplifiers The book also includes the discussion of the Barkhausen criterion for oscillations and the detailed analysis of various oscillator circuits including RC phase shift Wien bridge Hartley Colpitt s Clapp and crystal oscillators The book uses straightforward and lucid language to explain each topic The book provides the logical method of describing the various complicated issues and stepwise methods to make understanding easy The variety of solved examples is the feature of this book The book explains the subject s philosophy which makes understanding the concepts evident and makes the subject more interesting

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