

Section 7.4 - Determining Chemical Formulas

Calculation of Empirical Formulas - example

Let's Determine the empirical formula for a compound which is 54.09% Ca, 43.18% O, and 2.73% H

1) Divide each percent by that element's atomic weight.

$$\text{Ca} = 54.09/40 = 1.352$$

$$1.352/1.352 = 1$$

$$\text{O} = 43.18/16 = 2.699$$

$$2.699/1.352 = 2$$

$$\text{H} = 2.73/1 = 2.73$$

$$2.73/1.352 = 2$$

2) To get the answers to whole numbers, divide through by the smallest one.

This gives us $\rightarrow \text{CaO}_2\text{H}_2$ better yet $\rightarrow \text{Ca(OH)}_2$

Section Quiz Determining Chemical Formulas

Hock Leong Oon



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