

Colligative properties are properties that depend only upon the number of solute atoms, ions, or molecules in a solution and not on the nature of those atoms, ions or molecules. Freezing point depression and boiling point elevation are examples of colligative properties. Raoult discovered that the addition of solute particles causes the boiling point of a solution to be elevated and the freezing point to be depressed. It has been found that 1 mole of nonvolatile, nonionizing solute particles will raise the boiling point of 1.0 kg of water by 0.52 °C, to 100.52°C at 1 atm of pressure. The same concentration will lower the freezing point of 1.0 kg of water by 1.86°C, to -1.86°C. These two numbers apply only to water and are called the molal boiling point and freezing point constants for water. Other solvents have different constants.

PROPERTY	BOILING POINT ELEVATION	FREEZING POINT DEPRESSION
EQUATION	Boiling Point Elevation: A solution will boil at a higher temperature than the pure solvent. $\Delta T_b = i K_b m$ ΔT_b = change in boiling temp. in (°C) <i>i</i> = a unitless constant (called the van't Hoff factor) associated with the degree of dissociation of the solute in a solvent. K_b = molal boiling point elevation constant (°C/m) For H₂O: K_b = .52°C/m <i>m</i> = molality of the solution	Freezing Point Depression: A solution will freeze at a lower temperature than the pure solvent. $\Delta T_f = i K_f m$ ΔT_f = change in freezing temp. in (°C) <i>i</i> = a unitless constant (called the van't Hoff factor) associated with the degree of dissociation of the solute in a solvent. K_f = molal freezing point lowering constant (°C/m) For H₂O: K_f = 1.86°C/m <i>m</i> = molality of the solution
	<i>i</i> = 1 for substances that do not ionize 2 for substances which ionize into two ions like LiCl 3 for substances which ionize into three ions like CaCl ₂ ...etc. <i>i</i> may have to be calculated if a % of dissociation is given for a solute. If the solvent is not water you must find its K_b .	<i>i</i> = 1 for substances that do not ionize 2 for substances which ionize into two ions like LiCl 3 for substances which ionize into three ions like CaCl ₂ ...etc. <i>i</i> may have to be calculated if a % of dissociation is given for a solute. If the solvent is not water you must find its K_f .

EXAMPLE #1

If 90.0 g of nonionizing C₆H₁₂O₆, are dissolved in 255 g of H₂O, what will be the boiling point of the resulting solution?

First calculate the molality (*m*) of the solution:

$$m = \frac{92\text{g}}{\left(180 \frac{\text{g}}{\text{mole}}\right)(0.255 \text{ kg of solvent})} = 2.0 \text{ m}$$

Next apply the boiling point elevation equation:

$$[\textit{i} = 1 \text{ for a nonionizing solute.}]$$

$$\Delta T_b = i K_b m$$

$$\Delta T_b = (1)(0.52^\circ\text{C/m})(2.0 \text{ m})$$

$$\Delta T_b = 1.04^\circ\text{C}$$

$$\text{BP} = (\Delta T_b + 100.00^\circ\text{C})$$

$$\text{BP} = (1.04^\circ\text{C} + 100.00^\circ\text{C}) = \underline{101.04^\circ\text{C}}$$

Section Review 16 3 Colligative Properties Of Solutions

Donald A. McQuarrie, Stanley Gill



Section Review 16 3 Colligative Properties Of Solutions:

Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001

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Table of Contents Section Review 16 3 Colligative Properties Of Solutions

1. Understanding the eBook Section Review 16 3 Colligative Properties Of Solutions
 - The Rise of Digital Reading Section Review 16 3 Colligative Properties Of Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Section Review 16 3 Colligative Properties Of Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section Review 16 3 Colligative Properties Of Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section Review 16 3 Colligative Properties Of Solutions
 - Personalized Recommendations
 - Section Review 16 3 Colligative Properties Of Solutions User Reviews and Ratings

- Section Review 16 3 Colligative Properties Of Solutions and Bestseller Lists
- 5. Accessing Section Review 16 3 Colligative Properties Of Solutions Free and Paid eBooks
 - Section Review 16 3 Colligative Properties Of Solutions Public Domain eBooks
 - Section Review 16 3 Colligative Properties Of Solutions eBook Subscription Services
 - Section Review 16 3 Colligative Properties Of Solutions Budget-Friendly Options
- 6. Navigating Section Review 16 3 Colligative Properties Of Solutions eBook Formats
 - ePub, PDF, MOBI, and More
 - Section Review 16 3 Colligative Properties Of Solutions Compatibility with Devices
 - Section Review 16 3 Colligative Properties Of Solutions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section Review 16 3 Colligative Properties Of Solutions
 - Highlighting and Note-Taking Section Review 16 3 Colligative Properties Of Solutions
 - Interactive Elements Section Review 16 3 Colligative Properties Of Solutions
- 8. Staying Engaged with Section Review 16 3 Colligative Properties Of Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section Review 16 3 Colligative Properties Of Solutions
- 9. Balancing eBooks and Physical Books Section Review 16 3 Colligative Properties Of Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section Review 16 3 Colligative Properties Of Solutions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section Review 16 3 Colligative Properties Of Solutions
 - Setting Reading Goals Section Review 16 3 Colligative Properties Of Solutions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section Review 16 3 Colligative Properties Of Solutions
 - Fact-Checking eBook Content of Section Review 16 3 Colligative Properties Of Solutions
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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