

Chemistry I Study Guide Chapters 1-3

Chapter 1

Key Concepts

1. *What is a molecule?*
2. *What are the types of molecules?*
3. *What are the three types of bonds in molecules?*
4. *What are the four types of molecules?*
5. *What are the five types of molecules?*
6. *What are the six types of molecules?*
7. *What are the seven types of molecules?*
8. *What are the eight types of molecules?*
9. *What are the nine types of molecules?*
10. *What are the 10 types of molecules?*

Important Events

1. *In 1827, Friedrich Wöhler synthesized urea from carbons and hydrogens.*
2. *In 1828, Justus von Liebig isolated and identified nitrogen as a component of organic compounds.*
3. *In 1828, Michael Faraday discovered that electricity and magnetism are related.*
4. *In 1831, Friedrich Wöhler synthesized ammonia from nitrogen and hydrogen.*
5. *In 1835, Jean-Baptiste Dumas and Michel Eugène Chevreul identified sulfur as a component of organic compounds.*
6. *In 1839, Friedrich Wöhler synthesized glycerin from fats and water.*
7. *In 1841, John Dalton proposed that all elements are made of atoms of the same type.*
8. *In 1844, Lewis Paul and John Dalton independently proposed that the elements are arranged in a periodic table.*
9. *In 1845, Friedrich Wöhler synthesized urea from carbon and hydrogen.*
10. *In 1851, Louis Pasteur demonstrated that fermentation is caused by bacteria.*

Notable Figures

1. *Justus von Liebig*
2. *Friedrich Wöhler*
3. *Michael Faraday*
4. *Jean-Baptiste Dumas*
5. *John Dalton*
6. *Lewis Paul*
7. *Friedrich Wöhler*

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Chapter 1

Michael Rajnik



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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web bonus problem here is a balanced net ionic equation $3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$ the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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