

2.1

CHEMICAL REACTIONS

Section Review

Objectives

1. Measure enthalpy using a calorimeter.
2. Identify and predict chemical reactions using phase diagrams.
3. Apply the law of conservation of energy to chemical reactions.

Vocabulary

- | | | |
|------------------|---------------|----------------------------------|
| 1. endothermic | 2. exothermic | 3. enthalpy |
| 4. heat capacity | 5. joule | 6. law of conservation of energy |

Part A Completion

Use the information available to identify a chemical reaction, calculate, and write energy changes associated with the reaction using appropriate units and signs, identifying the system.

- | | |
|---|-----------|
| 1. Chemical change occurs when a substance _____. | 6. _____ |
| 2. Chemical change involves the transformation of a substance from _____. | 7. _____ |
| 3. Heat capacity is a _____ change. It is measured _____. | 8. _____ |
| 4. Heat _____ is the change in enthalpy of a substance _____. | 9. _____ |
| 5. _____ change is associated with heat and _____ change is associated _____. | 10. _____ |
| 6. Heat and work are _____ quantities. They have the _____ unit. | 11. _____ |
| 7. Heat is transferred from a hot object to a cold object _____. | 12. _____ |
| 8. Heat _____ is the amount of heat transferred _____. | 13. _____ |

Part B True-False

Indicate whether each statement is true or false. If false, explain why.

1. Heat capacity is a state function.
2. Heat capacity is a state function.
3. Heat capacity is a state function.
4. Heat capacity is a state function.
5. Heat capacity is a state function.
6. Heat capacity is a state function.
7. Heat capacity is a state function.
8. Heat capacity is a state function.
9. Heat capacity is a state function.
10. Heat capacity is a state function.

Section Review Neutralization Reactions Answers

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