

1. Given the nuclear equation:

The particle represented by X is

- A. ${}^6_3\text{Li}$
 B. ${}^4_2\text{He}$
 C. ${}^{10}_{10}\text{Be}$
 D. ${}^{10}_6\text{C}$
2. Which isotope will spontaneously decay and emit particles with a charge of +2?
 A. ${}^{55}_{26}\text{Fe}$
 B. ${}^{137}_{55}\text{Cs}$
 C. ${}^{208}_{81}\text{Au}$
 D. ${}^{226}_{88}\text{Fr}$
3. Which reaction is an example of natural transmutation?
 A. ${}^{238}_{94}\text{Pu} \rightarrow {}^{234}_{92}\text{U} + {}^4_2\text{He}$
 B. ${}^{27}_{13}\text{Al} + {}^4_2\text{He} \rightarrow {}^{30}_{15}\text{P} + {}^1_0\text{n}$
 C. ${}^{238}_{92}\text{U} + {}^1_0\text{n} \rightarrow {}^{235}_{92}\text{Pu} + 2 {}^0_{-1}\text{e}$
 D. ${}^{238}_{94}\text{Pu} + {}^1_0\text{n} \rightarrow {}^{140}_{54}\text{Ba} + {}^{94}_{38}\text{Sr} + 3 {}^1_0\text{n}$
4. What is the mass number of an alpha particle?
 A. 1
 B. 2
 C. 0
 D. 4
5. Which group of nuclear emissions is listed in order of increasing charge?
 A. alpha particle, beta particle, gamma radiation
 B. gamma radiation, alpha particle, beta particle
 C. positron, alpha particle, neutron
 D. neutron, positron, alpha particle
6. Which notation of a radioisotope is correctly paired with the notation of its emission particle?
 A. ${}^{37}_{20}\text{Ca}$ and ${}^4_2\text{He}$
 B. ${}^{238}_{92}\text{U}$ and ${}^0_{-1}\text{e}$
 C. ${}^{14}_7\text{N}$ and ${}^1_1\text{p}$
 D. ${}^3_1\text{H}$ and ${}^0_{-1}\text{e}$
7. Positrons are spontaneously emitted from the nuclei of
 A. potassium-37
 B. radium-226
 C. nitrogen-16
 D. thorium-232
8. Given the nuclear equation:

What particle is represented by X ?

- A. alpha
 B. beta
 C. neutron
 D. positron

9. Which nuclear emission has the greatest penetrating power?

- A. alpha particle
 B. beta particle
 C. gamma radiation
 D. positron

10. Which nuclear decay emission consists of energy, only?

- A. alpha particle
 B. beta particle
 C. gamma radiation
 D. positron

11. If
- $\frac{1}{4}$
- of an original sample of krypton-74 remains unchanged after 34.5 minutes, what is the half-life of krypton-74?

- A. 11.5 min
 B. 23.0 min
 C. 34.5 min
 D. 66.0 min

12. What is the half-life of sodium-25 if 1.00 gram of a 16.00-gram sample of sodium-25 remains unchanged after 23.7 seconds?

- A. 47.4 s
 B. 59.3 s
 C. 79.0 s
 D. 118 s

13. Atoms of one element are converted to atoms of another element through

- A. fermentation
 B. oxidation
 C. polymerization
 D. transmutation

14. Given the nuclear equation:

Which particle is represented by X ?

- A. ${}^4_2\text{He}$
 B. ${}^0_{-1}\text{e}$
 C. ${}^1_0\text{n}$
 D. ${}^0_{+1}\text{e}$

15. Which equation represents the radioactive decay of
- ${}^{226}_{88}\text{Ra}$
- ?

- A. ${}^{226}_{88}\text{Ra} \rightarrow {}^{222}_{86}\text{Rn} + {}^4_2\text{He}$
 B. ${}^{226}_{88}\text{Ra} \rightarrow {}^{226}_{89}\text{Ac} + {}^0_{-1}\text{e}$
 C. ${}^{226}_{88}\text{Ra} \rightarrow {}^{226}_{87}\text{Fr} + {}^0_{-1}\text{e}$
 D. ${}^{226}_{88}\text{Ra} \rightarrow {}^{222}_{88}\text{Ra} + {}^4_2\text{n}$

Nuclear Chemistry Regents Questions

L Cohen



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