

ANSWERS

Nuclear Chemistry Review Questions

1. Compare the mass, charge, and penetration power of alpha, beta and gamma radiation.

	Mass	Charge	Penetration
Alpha	4	+	Lowest
Beta	0	-	Middle
Gamma	0	0	Highest

2. In 5.49 seconds, 1.20g of argon-35 decays to leave only 0.15g. What is the half-life?

1.20g \rightarrow 0.60g \rightarrow 0.30g \rightarrow 0.15g , therefore it took 3 half-life periods.

$$5.49 \div 3 = 1.83 \text{ sec}$$

3. How many days does it take 16.0g of gold-198 to decay to 1.0g?

16.0g \rightarrow 8.0g \rightarrow 4.0g \rightarrow 2.0g \rightarrow 1.0g , therefore 4 half-life periods

From Ref Table N, 1 half-life = 2.69d

$$2.69\text{d} \times 4 = 10.76 \text{ days}$$

4. Why does an atom undergo radioactive decay?

Because it has an unstable nucleus

5. Compare and contrast fission and fusion. Why is fission used for electrical and fusion is not?

Fission=Division, the splitting of large unstable atoms(nucleus)

Fusion=joining together, small isotopes of H, creates way more energy

6. How is nuclear chemistry applied to medicine? What type of radioisotope must be used when it is injected into a person?

Radioisotopes are used to treat and diagnose illnesses

If injected they must have short half-lives and quickly removed from the body

7. Write out the nuclear decay of radium-226.



8. What is different about nuclear reactions when compared to a chemical reaction?

Nuclear reactions transform atoms into other atoms, chemical reactions involved the rearrangement of electrons

9. At what point do all isotopes of an element become radioactive?

Up to and including atomic # 83

Nuclear Chemistry Review Test Answers

Leo J. Malone



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Table of Contents Nuclear Chemistry Review Test Answers

1. Understanding the eBook Nuclear Chemistry Review Test Answers
 - The Rise of Digital Reading Nuclear Chemistry Review Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Nuclear Chemistry Review Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nuclear Chemistry Review Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nuclear Chemistry Review Test Answers
 - Personalized Recommendations
 - Nuclear Chemistry Review Test Answers User Reviews and Ratings
 - Nuclear Chemistry Review Test Answers and Bestseller Lists

5. Accessing Nuclear Chemistry Review Test Answers Free and Paid eBooks
 - Nuclear Chemistry Review Test Answers Public Domain eBooks
 - Nuclear Chemistry Review Test Answers eBook Subscription Services
 - Nuclear Chemistry Review Test Answers Budget-Friendly Options
6. Navigating Nuclear Chemistry Review Test Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Nuclear Chemistry Review Test Answers Compatibility with Devices
 - Nuclear Chemistry Review Test Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nuclear Chemistry Review Test Answers
 - Highlighting and Note-Taking Nuclear Chemistry Review Test Answers
 - Interactive Elements Nuclear Chemistry Review Test Answers
8. Staying Engaged with Nuclear Chemistry Review Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nuclear Chemistry Review Test Answers
9. Balancing eBooks and Physical Books Nuclear Chemistry Review Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nuclear Chemistry Review Test Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Nuclear Chemistry Review Test Answers
 - Setting Reading Goals Nuclear Chemistry Review Test Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Nuclear Chemistry Review Test Answers
 - Fact-Checking eBook Content of Nuclear Chemistry Review Test Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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