

Study Guide Answers

Chapter 21

Name: _____

Date: _____ Per: _____

- The splitting of the nucleus into lighter nuclei is called fission.
- When light mass nuclei combine to form a heavier, more stable nucleus this is called fusion.
- Write an equation for the alpha decay of bismuth-213. $^{213}_{83}\text{Bi} \rightarrow ^4_2\text{He} + ^{209}_{81}\text{Tl}$
- Write an equation for the beta decay of silver-106. $^{106}_{47}\text{Ag} \rightarrow ^0_{-1}\text{e} + ^{106}_{48}\text{Cd}$
- Describe the difference between beta decay and electron capture. In beta decay, an electron is emitted from the nucleus as a neutrons transforms into a proton. In electron capture (or K capture or beta capture) an electron is taken in from the electron cloud by the nucleus to convert a proton into a neutron.

- The mass of a lithium-7 atom* is 7.01600amu. Calculate it mass defect.

$$\begin{aligned}
 p^+ \text{ mass} &= 3 \times 1.007276 \text{ amu} &= 3.021828 \text{ amu} \\
 n^0 \text{ mass} &= 4 \times 1.008665 \text{ amu} &= 4.034660 \text{ amu} \\
 e^- \text{ mass} &= 7 \times 0.0005486 \text{ amu} &= 0.0038402 \text{ amu} \text{ (* mass is for atom, not nucleus)} \\
 \text{Theoretical Mass} &= 7.0581338 \text{ amu} \\
 \text{Mass defect} &= 7.0581338 \text{ amu} - 7.01600 \text{ amu} = \boxed{0.04213 \text{ amu}}
 \end{aligned}$$

Proton Mass 1.007276amu
Neutron Mass 1.008665amu
Electron Mass 0.0005486amu

- Calculate the nuclear binding energy of one lithium-6 atom. The actual measured atomic mass of lithium-6 is 6.015amu.

$$\begin{aligned}
 p^+ \text{ mass} &= 3 \times 1.007276 \text{ amu} &= 3.021828 \text{ amu} \\
 n^0 \text{ mass} &= 3 \times 1.008665 \text{ amu} &= 3.025995 \text{ amu} \\
 e^- \text{ mass} &= 7 \times 0.0005486 \text{ amu} &= 0.0038402 \text{ amu} \text{ (* mass is for atom, not nucleus)} \\
 \text{Theoretical Mass} &= 6.049683 \text{ amu} \\
 \text{Mass defect} &= 6.049683 \text{ amu} - 6.015 \text{ amu} = 0.034683 \text{ amu} \text{ (must be expressed in kg to use } E=mc^2\text{)}
 \end{aligned}$$

$$\frac{0.034683 \text{ amu}}{6.022 \times 10^{23} \text{ amu}} \times \frac{1 \text{ gram}}{1000 \text{ grams}} = 5.723812687 \times 10^{-29} \text{ kg}$$

$$E = mc^2 = 5.723812687 \times 10^{-29} \text{ kg} \cdot (3.00 \times 10^8 \text{ m/s})^2 = \boxed{5.15 \times 10^{-13} \text{ J}}$$

Since I used 3 sig figs for "c", I rounded the final answer to 3 sig figs.

- Oxygen has an unstable isotope O-17 that has a mass of 17.00454. Calculate the binding energy of the oxygen nucleus in MeV.

$$\begin{aligned}
 p^+ \text{ mass} &= 8 \times 1.007276 \text{ amu} &= 8.058208 \text{ amu} \\
 n^0 \text{ mass} &= 9 \times 1.008665 \text{ amu} &= 9.077985 \text{ amu} \\
 e^- \text{ mass} &= 8 \times 0.0005486 \text{ amu} &= 0.0043888 \text{ amu} \text{ (* mass is for atom, not nucleus)} \\
 \text{Theoretical Mass} &= 17.1405818 \text{ amu} \\
 \text{Mass defect} &= 17.1405818 \text{ amu} - 17.00454 = 0.1360418 \text{ amu} \text{ (must be expressed in kg to use } E=mc^2\text{)}
 \end{aligned}$$

$$\frac{0.1360418 \text{ amu}}{6.022 \times 10^{23} \text{ amu}} \times \frac{1 \text{ gram}}{1000 \text{ grams}} = 2.25908004 \times 10^{-28} \text{ kg}$$

$$E = mc^2 = 2.25908004 \times 10^{-28} \text{ kg} \cdot (3.00 \times 10^8 \text{ m/s})^2 = 2.03382036 \times 10^{-11} \text{ J}$$

1eV = 1.60 x 10⁻¹⁹ joules (*I give this value to you in the text)

$$\frac{2.03382036 \times 10^{-11} \text{ J}}{1.60 \times 10^{-19} \text{ J}} \times \frac{1 \text{ MeV}}{1000000 \text{ eV}} = \boxed{127 \text{ MeV}}$$

Since I used 3 sig figs for "c" and the eV conversion, I rounded the final answer to 3 sig figs.

Nuclear Chemistry Packet Answers Chapter 21

B Lingard



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