

Unit 2: Atoms and Elements
AP Chemistry
The Nuclear Atom

ALL of the answers to this worksheet can be **logically** figured out by looking at the **Schematic Diagrams for Various Atoms**, the **Periodic Table**, and **discussing** with your partners. All of the information you need is here somewhere. Determine each answer and be able to give convincing reasons for each answer. Good luck.

1. How many **protons** are found in ^{12}C ? **6** ^{13}C ? **6** ^{14}C ? **6**
2. How many **neutrons** are found in ^{12}C ? **6** ^{13}C ? **7** ^{14}C ? **7**
3. How many **electrons** are found in ^{12}C ? **6** ^{13}C ? **6** ^{14}C ? **7**
4. Based on the model,
 - a) what do all carbon atoms (and ions) have in common?
Same number and types of protons.
 - b) what do all hydrogen atoms (and ions) have in common?
Same number and types of protons.
5. What is the significance of the atomic number, Z , above each atomic symbol in the periodic chart?
It tells us the number of protons located in the nucleus of an atom of that particular element.
6. What do all nickel (Ni) atoms have in common?
Same number and types of protons.
7. How is the mass number, A , (left-hand superscript next to the atomic symbol) determined?
By adding the protons and neutrons of an atom together.
8. What structural feature is different in isotopes of a particular element?
The number of neutrons that are present.
9.
 - a) What feature distinguishes a neutral atom from an ion?
The number of electrons that are present.
 - b) How is the charge on an ion determined?
By subtracting the number of electrons from the number of protons in the atom.
10. Where is most of the mass of an atom, within the nucleus or outside of the nucleus? Explain your reasoning.
Within the nucleus since that is where the protons and neutrons are located.

11. Complete the following table:

Isotope	Atomic Number Z	Mass Number A	Number of electrons
^{31}P	15	31	15
^{18}O	8	18	8
$^{39}\text{Y}^{+}$	19	39	18
$^{58}\text{Ni}^{2+}$	28	58	26

Nuclear Power Chapter 22 Number 1 Answer Key

D Keegan



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