

Unit 3 Study Guide: Atomic Structure and Nuclear Energy

Section 4.1 Early Ideas About Matter

For each statement below, write *true* or *false*.

- _____ 1. Ancient philosophers regularly performed controlled experiments.
- _____ 2. Philosophers formulated explanations about the nature of matter based on their own experiences.
- _____ 3. Both Democritus and Dalton suggested that matter is made up of atoms.
- _____ 4. Dalton's atomic theory stated that atoms separate, combine, or rearrange in chemical reactions.
- _____ 5. Dalton's atomic theory stated that matter is mostly empty space.
- _____ 6. Dalton was correct in thinking that atoms could not be divided into smaller particles.
- _____ 7. Dalton's atomic theory stated that atoms of different elements combine in simple whole-number ratios to form compounds.
- _____ 8. Dalton thought that all atoms of a specific element have the same mass.
- _____ 9. Democritus proposed that atoms are held together by chemical bonds, but no one believed him.
- _____ 10. Dalton's atomic theory was based on careful measurements and extensive research.

Section 4.2 Defining the Atom

For each item in Column A, write the letter of the matching item in Column B.

Column A

- _____ 1. Proposed the nuclear atomic model
- _____ 2. Determined the mass-to-charge ratio of an electron
- _____ 3. Calculated the mass of an electron

Column B

- a. Thomson
- b. Millikan
- c. Rutherford

Draw and label a diagram of each atomic model.

4. plum pudding model

5. nuclear atomic model

Nuclear Unit Study Guide

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