

Chapter 19 – Nuclear Chemistry

This packet contains background and explanatory information on Nuclear Chemistry in addition to practice handouts. The information should be read and used to complete all of the practice included. I will be collecting this packet and giving points for the completion of the contents. You may want to consult your textbook (Chapter 19) or online resources for addition explanations and examples if needed.

Chapter Objectives: The student will (be able to)...

- Compare and contrast **alpha**, beta, and gamma radiation.
- Identify alpha decay, beta decay, fission, and fusion reactions.
- Balance nuclear equations.
- Solve half-life calculations.

Part I: Radioactivity - What is it? <http://www.darvill.clara.net/nucrad/index.htm>

All substance are made of **atoms**. These have **electrons** (e) around the outside, and a **nucleus** in the middle. The nucleus consists of **protons** (p) and **neutrons** (n), and is extremely small. (Atoms are almost entirely made of empty space!)

In some types of atom, the **nucleus** is **unstable**, and will **decay** into a more stable atom. This **radioactive decay** is completely **spontaneous**. The energy that is released from the nucleus of the atom is **radiation**.

When an unstable nucleus **decays**, there are **three** ways that it can do so. It may give out:-

- an **alpha particle** (we use the symbol α)
- a **beta particle** (symbol β)
- a **gamma ray** (symbol γ)

Many radioactive substances **emit** α particles and β particles as well as γ rays. In fact, you won't find a pure γ source: anything that gives off γ rays will also give off α and/or β too.

A. Alpha Particles – More Information

- Alpha particles are made of **2 protons** and **2 neutrons**.
- This means that they have a **charge** of **+2**, and a **mass** of **4** (the mass is measured in "atomic mass units", where each proton & neutron = 1)

We can write them as ${}^4_2\alpha$, or, because they're the same

as a helium nucleus, ${}^4_2\text{He}$.

- Alpha particles are relatively **slow** and **heavy**.
- They have a **low penetrating power** - you can stop them with just a **sheet of paper**.
- Because they have a large charge, alpha particles **ionize** (pull electrons from) other atoms strongly.



Alpha particle:
2 protons
& 2 neutrons

Nuclear Chemistry Packet Answers

NARAYAN CHANGDER



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Guide to U. S. Government Publications ,1980 **Excel With Objective Questions In Chemistry** Prof. S. K. Khanna,Dr. N. K. Verma,Dr. B. Kapila,2006 Excel With Objective Questions In Physical Chemistry S.K. Khanna; N.K. Verma; B. Kapila, **Resources in Education** ,1994-04 *Chemistry Problems* Joseph F. Castka,1962 **Guide to U.S Government Publications** Gale Group,2000-02 This highly respected single volume resource catalogs all important series periodicals and reference tools published by the federal government each year including annual reports general publications federal laws state laws regulations rules and instructions press releases and more **Nuclear Science Abstracts** ,1972
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novel developments to meet the current challenges in the field Further this monograph responds to a need for a systematic comparative treatise on nonadiabatic theories of quantum molecular dynamics which are of considerably higher complexity than the more traditional adiabatic approaches and are steadily gaining in importance This volume addresses a broad readership ranging from physics or chemistry graduate students to specialists in the field of theoretical quantum dynamics

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Vocational Division Bulletin ,1962 **Resources in Education ,1993** **Heinemann Physics for CXC** Norman Lambert,Natasha Lewis dos Santos,Tricia A. Samuel,2000 Heinemann Physics for CXC is a lively accessible textbook written by Norman Lambert the well repsected author and teacher and experienced teachers Natasha Lewis dos Santos and Tricia A Samuel The authors have drawn on their many years of teaching **Vocational Education Bulletin** United States. Division

of Vocational Education,1961 **Tunnelling in Molecules** Johannes Kästner,Sebastian Kozuch,2020-09-22 Quantum tunnelling is one of the strangest phenomena in chemistry where we see the wave nature of atoms acting in impossible ways By letting molecules pass through the kinetic barrier instead of over it this effect can lead to chemical reactions even close to the absolute zero to atypical spectroscopic observations to bizarre selectivity or to colossal isotopic effects Quantum mechanical tunnelling observations might be infrequent in chemistry but it permeates through all its disciplines producing remarkable chemical outcomes For that reason the 21st century has seen a great increase in theoretical and experimental findings involving molecular tunnelling effects as well as in novel techniques that permit their accurate predictions and analysis Including experimental computational and theoretical chapters from the physical and organic to the biochemistry fields from the applied to the academic arenas this new book provides a broad and conceptual perspective on tunnelling reactions and how to study them Quantum Tunnelling in Molecules is the obligatory stop for both the specialist and those new to this world

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