

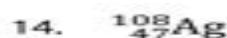
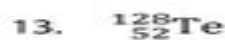
Unit 3 Study Guide: Atomic Structure and Nuclear Chemistry

Section 1: Atomic Structure

Define the following terms.

1. atom
2. Why is an atom neutral?
3. electron cloud
4. electrons
5. What determines which element an atom is?
6. neutrons
7. Which equals one amu—the mass of one atom of carbon-12, or 1/12 the mass of one atom of carbon-12?
8. atomic number
9. mass number
10. hyphen notation

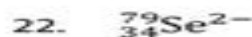
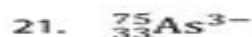
Give the number of protons, neutrons, and electrons for the following:



Section 2: Ions and Isotopes

15. Describe how cations are formed.
16. Describe how anions are formed.
17. What type of element generally forms cations?
18. How are isotopes of the same element alike and how are they different?

Give the number of protons, neutrons, and electrons for the following:



Section 3: Average Atomic Mass

23. Chlorine has two naturally occurring isotopes, Cl-35 and Cl-37. Using the periodic table, determine which of these two isotopes of chlorine is more abundant.

Nuclear Chemistry Section 1 Answers

Heather Hattori, Richard H. Langley



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Chemistry Jeopardy Brian Pressley, 2002 A quiz game designed to reinforce key chemistry concepts ideas and facts

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complexity of the numerous numerical problems found in most parts of the subject This text is crafted to provide a clear and accessible pathway to overcoming this challenge by assisting students especially novices or those with minimal knowledge of the subject in performing chemistry calculations The content covers fundamental calculations crucial to understanding the principles of chemistry making it an invaluable tool for students aiming to excel in their studies Key features Designed with a student friendly approach including detailed explanation of chemical concepts underlying each type of calculation step by step explanations alternative methods for solving problems numerous practice exercises answers to practice exercises and appendices The book is tailored to suit various curricula ensuring relevance for a diverse audience Encompasses a wide range of calculations offering students a thorough understanding of essential chemistry concepts Serves as an excellent resource for exam preparation and equips students with skills applicable to future scientific endeavours Employs straightforward language to ensure ease of understanding for beginners Uses IUPAC conventions underscoring the universal nature of chemistry

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The Burning Answer Keith Barnham,2015-05-15 Our society faces a choice We could be enjoying a sustainable lifestyle but we have chosen not to In three generations we have consumed half the oil produced by photosynthesis over eight million generations In two generations we have used half our uranium resources With threats from global warming oil depletion and nuclear disaster we are running out of options Solar power as Keith Barnham says is our necessary solution Barnham explains that the roots of solar energy lie in a little known equation $E = hf$ an equation which was coincidentally celebrated and explained to the world by Einstein in the same year he discovered $E = mc^2$ He alleges that the former equation has been overlooked in favor of the latter much to our detriment and Barnham is here to offer us a solution We can still turn things around and solar energy is the key In this provocative inspiring passionately argued book Keith Barnham outlines actions that any one and all of us can take to make an impact now and on future generations The Burning Answer is a solar manifesto for the new climate aware generation and a must read for climate change skeptics

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