

Unit 2: Atoms and Elements  
AP Chemistry  
The Nuclear Atom

ALL of the answers to this worksheet can be **logically** figured out by looking at the **Schematic Diagrams for Various Atoms**, the **Periodic Table**, and **discussing** with your partners. All of the information you need is here somewhere. Determine each answer and be able to give convincing reasons for each answer. Good luck.

1. How many **protons** are found in  $^{12}\text{C}$ ? **6**       $^{13}\text{C}$ ? **6**       $^{14}\text{C}$ ? **6**
2. How many **neutrons** are found in  $^{12}\text{C}$ ? **6**       $^{13}\text{C}$ ? **7**       $^{14}\text{C}$ ? **7**
3. How many **electrons** are found in  $^{12}\text{C}$ ? **6**       $^{13}\text{C}$ ? **6**       $^{14}\text{C}$ ? **7**
4. Based on the model,
  - a) what do all carbon atoms (and ions) have in common?  
**Same number and types of protons.**
  - b) what do all hydrogen atoms (and ions) have in common?  
**Same number and types of protons.**
5. What is the significance of the atomic number,  $Z$ , above each atomic symbol in the periodic chart?  
**It tells us the number of protons located in the nucleus of an atom of that particular element.**
6. What do all nickel ( $\text{Ni}$ ) atoms have in common?  
**Same number and types of protons.**
7. How is the mass number,  $A$ , (left-hand superscript next to the atomic symbol) determined?  
**By adding the protons and neutrons of an atom together.**
8. What structural feature is different in isotopes of a particular element?  
**The number of neutrons that are present.**
9.
  - a) What feature distinguishes a neutral atom from an ion?  
**The number of electrons that are present.**
  - b) How is the charge on an ion determined?  
**By subtracting the number of electrons from the number of protons in the atom.**
10. Where is most of the mass of an atom, within the nucleus or outside of the nucleus? Explain your reasoning.  
**Within the nucleus since that is where the protons and neutrons are located.**
11. Complete the following table:

Isotope	Atomic Number $Z$	Mass Number $A$	Number of electrons
$^{31}\text{P}$	15	31	15
$^{18}\text{O}$	8	18	8
$^{39}\text{Y}^{+}$	19	39	18
$^{58}\text{Ni}^{2+}$	28	58	26

# Nuclear Chemistry Worksheet Answer Key

**Ensheng Dong**



## **Nuclear Chemistry Worksheet Answer Key:**

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