

# **Molecular Chemistry** (part of a curriculum)

Model 1: write symbols for three isotopes of carbon



## **Critical Thinking Questions**

- How many protons are in carbon-12?  $6$  How many neutrons are in carbon-12?  $6$
- How many protons are in carbon-13?  $6$  How many neutrons are in carbon-13?  $7$
- How many protons are in carbon-14?  $6$  How many neutrons are in carbon-14?  $8$
- What is similar about the three and what different about them?

Students must think critically, consider, assess, evaluate, and synthesize information. They are asked to think about the isotopes of carbon and their effects on the environment.

- What does the subscript in chemical symbols mean? What does the superscript mean? What does the symbol mean?

Students must think about the meaning of the symbols and the meaning of the subscript and superscript in chemical symbols.

A student responds to a question in the classroom or on a worksheet as shown. This is not necessarily a correct answer to the question, but it shows the student's understanding of the question.

What are three types of carbon isotopes? Radioactive, stable, and unstable. Radioactive isotopes of carbon are used in dating. Stable isotopes are used in medicine. Unstable isotopes are used in industry.

Radioactive isotopes are unstable, meaning that their nuclei break apart. Stable isotopes are stable, meaning that their nuclei do not break apart.

Table 1: Carbon isotopes and radiocarbon dating used in molecular chemistry

Isotope of Carbon	Symbol	Atomic Number	Mass Number	Relative Abundance	Stability	Uses
Carbon-12	$^{12}_6\text{C}$	6	12	98.9%	Stable	Radioactive dating, medical imaging
Carbon-13	$^{13}_6\text{C}$	6	13	1.1%	Stable	Radioactive dating, medical imaging
Carbon-14	$^{14}_6\text{C}$	6	14	0.011%	Radioactive	Radioactive dating, medical imaging
Carbon-15	$^{15}_6\text{C}$	6	15	0.0001%	Radioactive	Radioactive dating, medical imaging
Carbon-16	$^{16}_6\text{C}$	6	16	0.0001%	Radioactive	Radioactive dating, medical imaging

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**Jim Schneider**



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