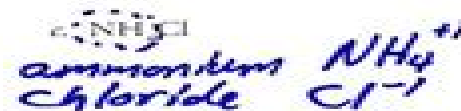
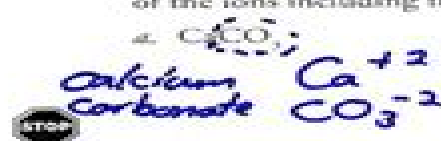


15. Bromine forms polyatomic ions with structures similar to those of chlorine. Using the chlorine family of polyatomic ions as a model, predict the name of the BrO_4^- ion.

ClO_4^- is perchlorate, therefore, BrO_4^- is perbromate.

16. Identify the polyatomic ion in each of these ionic compounds. Write out the name and formula of the ions including their charges.



Model 3 – Ternary Ionic Compounds

Compound Name	Ion Symbols and Charges		Chemical Formula
Ammonium phosphate	NH_4^{+1}	PO_4^{3-}	$(\text{NH}_4)_3\text{PO}_4$
Barium nitrite	Ba^{2+}	NO_2^{-1}	$\text{Ba}(\text{NO}_2)_2$
Ammonium sulfate	NH_4^{+1}	SO_4^{2-}	$(\text{NH}_4)_2\text{SO}_4$
Aluminum carbonate	Al^{3+}	CO_3^{2-}	$\text{Al}_2(\text{CO}_3)_3$
Iron(III) hydroxide	Fe^{3+}	OH^{-1}	$\text{Fe}(\text{OH})_3$
Potassium nitrate	K^{+1}	NO_3^{-1}	KNO_3

17. How are ternary ionic compounds in Model 3 different from binary ionic compounds (NaCl , MgO , CaBr_2 , etc.) that you've seen previously? *Hint:* Consider the meaning of the word "binary."

Ternary ionic compounds contain more than two elements.
 Binary ionic compounds contain only two elements.

18. Consider the compound iron(III) hydroxide in Model 3.

- a. How many hydroxide ions (OH^{-1}) are combined with an iron(III) ion (Fe^{3+})?

There are three hydroxide (OH^{-1}) for every one iron(III) (Fe^{3+}) ion.

- b. Is your answer to part a the only combination of iron(III) and hydroxide that should exist in nature? Explain.

Yes. The ratio of ~~three~~ ^{one} Fe^{3+} and three OH^{-1} is the only ratio that can balance the charge of the two ions.

19. Consider the compound barium nitrite in Model 3.

- a. What does the subscripted "2" inside the parentheses of the chemical formula tell you about the compound?

The "2" inside of the parentheses indicates that there are two oxygen atoms in the nitrite (NO_2^{-1}) polyatomic ion.



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NARAYAN CHANGDER



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 - The Rise of Digital Reading Nuclear Chemistry Worksheet K Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Nuclear Chemistry Worksheet K Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nuclear Chemistry Worksheet K Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nuclear Chemistry Worksheet K Answer Key
 - Personalized Recommendations
 - Nuclear Chemistry Worksheet K Answer Key User Reviews and Ratings
 - Nuclear Chemistry Worksheet K Answer Key and Bestseller Lists
5. Accessing Nuclear Chemistry Worksheet K Answer Key Free and Paid eBooks
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 - Nuclear Chemistry Worksheet K Answer Key Budget-Friendly Options
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- Nuclear Chemistry Worksheet K Answer Key Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Nuclear Chemistry Worksheet K Answer Key
 - Highlighting and Note-Taking Nuclear Chemistry Worksheet K Answer Key
 - Interactive Elements Nuclear Chemistry Worksheet K Answer Key
- 8. Staying Engaged with Nuclear Chemistry Worksheet K Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nuclear Chemistry Worksheet K Answer Key
- 9. Balancing eBooks and Physical Books Nuclear Chemistry Worksheet K Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nuclear Chemistry Worksheet K Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Nuclear Chemistry Worksheet K Answer Key
 - Setting Reading Goals Nuclear Chemistry Worksheet K Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Nuclear Chemistry Worksheet K Answer Key
 - Fact-Checking eBook Content of Nuclear Chemistry Worksheet K Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
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