# **Nuclear Decay**

#### Types of Radiation

Complete the table below to compare the properties of alpha, beta, and gamma radiation.

Name	Alpha	Beta	Gamma
Greek letter	a	β-	γ
Symbol	He	_5c	γ
Composition	2 protons, 2 neutrons	an electron	radiation (light)
Charge	O	-1	0
Stopped by	paper, skin, clothing	aluminum	lead

#### **Nuclear Decay Reactions**

In problems 2 - 5, fill in the missing particle for each nuclear decay reaction. Then, match each equation to the appropriate type of nuclear decay.

A. alpha emission

B. beta emission
C. electron capture

$$+_{-1}^{0}e \rightarrow_{\kappa_0}^{157}Nd$$

In problems 6 and 7, write a balanced equation for each nuclear decay reaction.

Decay of radium-223 by alpha (α) emission.

220 as Ra - 4 He + 230 km

Decay of iodine-126 by electron (β) emission.

128 al - 0 te + 128 a Xe

S. How are the atomic number and mass number of a nuclide affected by Beta Decay?

The mass number does not change but the atomic number increases by I

9. How are the atomic number and mass number of a nuclide affected by the emission of gamma radiation?

The atomic number and mass number do not change

10. Which has a slightly greater mass in a nuclear reaction, the reactants or the products? Why? It depends on the type of reaction. A faction reaction will have greater mass in the products because it is combining two nuclei. A faction reaction will not change mass.

# **Nuclear Decay Radioactive Worksheet Answer Key**

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