Unit 5, Activity 5

Balancing Redox Reactions using half-reaction method worksheet

Balance the following equations using the half-reaction method. Identify the reactant being oxidized and reduced.

a.
$$Cu^{2+} + I^{1-} \rightarrow Cu + I_3^{1-}$$
 (neutral)

b.
$$MnO_4^{1-} + Ag \rightarrow Mn^{2+} + Ag^{1+}$$
 (acidic)

c.
$$Hg + NO_3^{1-} + Cl^{1-} \rightarrow HgCl_4^{2-} + NO_2$$
 (acidic)

d.
$$AsH_3 + Zn^{2+} \rightarrow H_3AsO_4 + Zn$$
 (acidic)

e.
$$CN^{1-} + MnO_4^{1-} \rightarrow CNO^{1-} + MnO_2$$
 (basic)

g.
$$ClO^{1-} + CrO_2^{1-} \rightarrow CrO_4^{2-} + Cl_2$$
 (basic)

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