

REDOX REACTION

Solved Examples

Example - 6 :



Base Acid

Sol. Valency factor of base = 1

Here, two molecule of NaOH replaced 2H^+ ion from the H_2SO_4 . Therefore, each molecule of NaOH replaced only one H^+ ion of acid, so v.f. = 1.

○ v. f. for acid is the number of OH^- replaced from the base by each molecule of acid

Solved Examples

Example - 7 :



Base Acid

Sol. Valency factor of acid = 1

Here, one of molecule of H_2SO_4 replaced one OH^- from NaOH. Therefore, valency factor for H_2SO_4 is one

$$\therefore \text{Eq. wt. of } \text{H}_2\text{SO}_4 = \frac{\text{Mol. wt.}}{1}$$

○ Salts :

(a) In non-reacting condition

○ Valency factor = Total number of positive charge or negative charge present in the compound.

Solved Examples

Example - 8 :

		Na_2CO_3	$\text{Fe}_2(\text{SO}_4)_3$	$\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$
Sol.	v.f.	2	$2 \times 3 = 6$	2
	Eq. wt.	$\frac{M}{2}$	$\frac{M}{6}$	$\frac{M}{2}$

(b) In reacting condition

Solved Examples

Example - 9 :



Base Acid

Sol. It is an acid base reaction, therefore valency factor for Na_2CO_3 is one while in non-reacting condition, it will be two.

(c) Equivalent weight of oxidising / reducing agents in a redox reaction

In case of redox change, v.f. = Total change in oxidation number per molecule.

Redox Reactions Chemistry Study Guide Answers

David Kirk



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Table of Contents **Redox Reactions Chemistry Study Guide Answers**

1. Understanding the eBook Redox Reactions Chemistry Study Guide Answers
 - The Rise of Digital Reading Redox Reactions Chemistry Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Redox Reactions Chemistry Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Redox Reactions Chemistry Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Redox Reactions Chemistry Study Guide Answers
 - Personalized Recommendations
 - Redox Reactions Chemistry Study Guide Answers User Reviews and Ratings
 - Redox Reactions Chemistry Study Guide Answers and Bestseller Lists

5. Accessing Redox Reactions Chemistry Study Guide Answers Free and Paid eBooks
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 - Redox Reactions Chemistry Study Guide Answers eBook Subscription Services
 - Redox Reactions Chemistry Study Guide Answers Budget-Friendly Options
6. Navigating Redox Reactions Chemistry Study Guide Answers eBook Formats
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 - Redox Reactions Chemistry Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Redox Reactions Chemistry Study Guide Answers
 - Highlighting and Note-Taking Redox Reactions Chemistry Study Guide Answers
 - Interactive Elements Redox Reactions Chemistry Study Guide Answers
8. Staying Engaged with Redox Reactions Chemistry Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Redox Reactions Chemistry Study Guide Answers
9. Balancing eBooks and Physical Books Redox Reactions Chemistry Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Redox Reactions Chemistry Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Redox Reactions Chemistry Study Guide Answers
 - Setting Reading Goals Redox Reactions Chemistry Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Redox Reactions Chemistry Study Guide Answers
 - Fact-Checking eBook Content of Redox Reactions Chemistry Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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