

Relative Mass and the Mole

How can atoms be counted using a balance?

Why?

Consider the following equation for a chemical reaction: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

This can be interpreted as two molecules of hydrogen and one molecule of oxygen combining to form two water molecules. But how often do chemists limit their reactions to one or two molecules? Usually a reaction is done with an unimaginable number of molecules. How then do chemists know they have the right mix? The molecules need to be quickly counted! How do we count molecules? The answer is the unit called the **mole**. This activity will start by considering two egg farmers (a chicken farmer and a quail farmer). They produce such large numbers of eggs that they can't count them all individually, so they count in dozens of eggs in some cases, while in other cases they use mass. Weighing is often easier than counting!

Model 1 – Eggs

Chicken		Quail		Ratio of numbers of eggs	Ratio of masses of eggs
Number of eggs in the sample	Mass of the sample	Number of eggs in the sample	Mass of the sample		
1	37.44 g	1	2.34 g	1 : 1	16 : 1
10	374.4g	10	23.4g	1:1	16:1
438	16394	438	1.023g	1:1	16:1
1 dozen	449.3g	1 dozen	28.08g	1:1	16:1
1 million	37.	1 million	2.340,000 g	1:1	16:1

1. Consider the data in Model 1.

a. What is the mass of a standard chicken egg?

37.44g

b. What is the mass of a standard quail egg?

2.34g

c. Show mathematically how the 16:1 ratio of masses was calculated in the last column of Model 1.

$\frac{37.44}{2.34} = 16$

2. Use a calculator to complete the table in Model 1. Divide the work among group members. Reduce all ratios to the lowest whole numbers possible.

Relative Mass And The Mole Key

Yijin Wang



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Table of Contents Relative Mass And The Mole Key

1. Understanding the eBook Relative Mass And The Mole Key
 - The Rise of Digital Reading Relative Mass And The Mole Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Relative Mass And The Mole Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Relative Mass And The Mole Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Relative Mass And The Mole Key
 - Personalized Recommendations
 - Relative Mass And The Mole Key User Reviews and Ratings
 - Relative Mass And The Mole Key and Bestseller Lists
5. Accessing Relative Mass And The Mole Key Free and Paid eBooks
 - Relative Mass And The Mole Key Public Domain eBooks

- Relative Mass And The Mole Key eBook Subscription Services
- Relative Mass And The Mole Key Budget-Friendly Options
- 6. Navigating Relative Mass And The Mole Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Relative Mass And The Mole Key Compatibility with Devices
 - Relative Mass And The Mole Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Relative Mass And The Mole Key
 - Highlighting and Note-Taking Relative Mass And The Mole Key
 - Interactive Elements Relative Mass And The Mole Key
- 8. Staying Engaged with Relative Mass And The Mole Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Relative Mass And The Mole Key
- 9. Balancing eBooks and Physical Books Relative Mass And The Mole Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Relative Mass And The Mole Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Relative Mass And The Mole Key
 - Setting Reading Goals Relative Mass And The Mole Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Relative Mass And The Mole Key
 - Fact-Checking eBook Content of Relative Mass And The Mole Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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