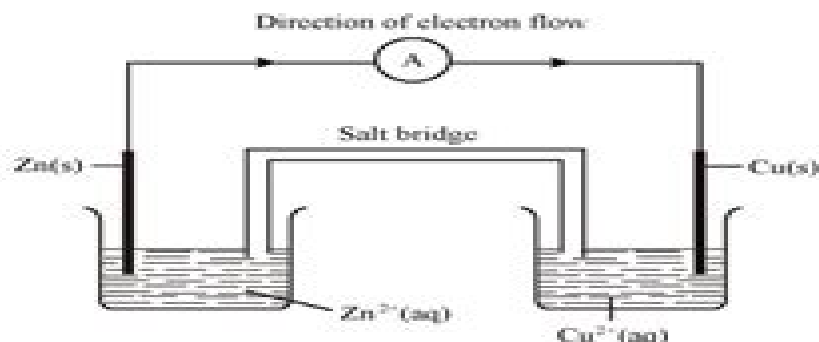


Redox Review – IB Short Questions

Question: 1



- (a) The apparatus shown above may be used to carry out a redox reaction.
- State the function of the salt bridge. 1
 - Write a half-equation for the oxidation reaction. 1
 - The above reactions are carried out under *standard conditions*. State what the standard conditions are for the cell. 2
 - Using the Data Booklet, calculate the cell potential for the above cell. 2
 - State and explain what happens to the concentration of the copper(II) ions when the cell is producing an electric current. 2
 - State **two** observations that could be made if the zinc rod were placed in a solution of copper(II) ions. 2
- (b) The standard electrode potentials for three electrode systems are given below.
- | | |
|--|-------------------------------|
| $\text{Ti}^{3+}(\text{aq}) + \text{e}^{-} \rightarrow \text{Ti}^{2+}(\text{aq})$ | $E^{\circ} = -0.37 \text{ V}$ |
| $\text{Fe}^{3+}(\text{aq}) + \text{e}^{-} \rightarrow \text{Fe}^{2+}(\text{aq})$ | $E^{\circ} = +0.77 \text{ V}$ |
| $\text{Ce}^{4+}(\text{aq}) + \text{e}^{-} \rightarrow \text{Ce}^{3+}(\text{aq})$ | $E^{\circ} = +1.45 \text{ V}$ |
- Using the data above, deduce which species is the best reducing agent, giving a reason in terms of electrons for your answer. 2
 - Write an equation, including state symbols, for the overall reaction with the greatest cell potential. 2

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