CHAPTER-8 REDOX REACTIONS

oxidation	reduction
Addition of oxygen	 Removal of oxygen
Removal of hydrogen	Addition of hydrogen
Addition of an electronegative element	 Removal of an electronegative element
Removal of an electropositive element	 Addition of an electropositive element
5. Loss of electron	Gain of electron

Oxidation number denotes theoxidation state of an element in a compound ascertained according to a set of rules formulated on the basis that electron in a covalent bond belongsentirely to more electronegative element.

Calculation of oxidation number-

- O. S. of all the elements in their elemental form (in standard state) is taken as zero O. S. of elements in Cl₂, F₂, O₂, P₄, O₃, Fe(s), H₂, N₂, C(graphite) is zero.
- Common O. S. of elements of group one (1st) is one. Common O. S. of elements of group two (2nd) is two.
- For ions composed of only one atom, theoxidation number is equal to the chargeon the ion.
- The oxidation number of oxygen in most compounds is −2. While in peroxides (e.g., H₂O₂, Na₂O₂), eachoxygen atom is assigned an oxidationnumber of −1, in superoxides (e.g., KO₂,RbO₂) each oxygen atom is assigned anoxidation number of −(½).
- In oxygendifluoride (OF₂) and dioxygendifluoride (O₂F₂), the oxygen is assigned an oxidation number of +2 and +1, respectively.
- The oxidation number of hydrogen is +1 but in metal hydride its oxidation no. is-1.
- In all its compounds, fluorine has anoxidation number of −1.
- The algebraic sum of the oxidation number of all the atoms in a compound must be zero.

Redox Reactions Study Guide Answer Key

Susan McMurry, John McMurry

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