

### Pre-Lab Assignment: Redox Titration of Bleach

A student attempting to determine the mass percent of iron in a sample of ore reacts the ore with strong acid to convert the Fe to  $\text{Fe}^{2+}$ . Then the  $\text{Fe}^{2+}$  is titrated with  $\text{MnO}_4^{1-}$  according to the following balanced equation:



The student used 1.1081 grams of the ore and needed 39.32 mL of 0.03190 M  $\text{KMnO}_4$  (the source of the  $\text{MnO}_4^{1-}$ ) to complete the titration.

1. How many moles of permanganate ion,  $\text{MnO}_4^{1-}$ , did the student add to the ore?
2. What mass of iron was found to be present in the ore?
3. Determine the mass percent of the iron in the ore.

# Redox Titration Pre Lab Questions Answers

**Rosina Ehmann**



## **Redox Titration Pre Lab Questions Answers:**

*Chemistry* Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02      *Newer Redox Titrants* A. Berka, J. Vulterin, J. Zýka, 2013-10-22 International Series of Monographs in Analytical Chemistry Volume 22

Newer Redox Titrants focuses on the processes reactions methodologies and approaches involved in the study of redox titrants The publication first offers information on potassium permanganate in alkaline solution and compounds of trivalent manganese including standard solutions indicator and review of determinations The text then ponders on compounds of trivalent copper and potassium hexacyanoferrate The book ponders on hypohalites hypochlorite and hypobromite chloramine T and bromine as well as standard solutions indicator and review of determinations The publication also takes a look at iodine monochloride periodic acid and its salts lead IV acetate compounds of pentavalent vanadium and iron III salts The compounds of trivalent cobalt hydrogen peroxide chromium II salts tin II chloride sodium arsenite and compounds of monovalent copper are also elaborated The publication is a reliable reference for readers interested in newer redox titrants

**Laboratory Manual for Principles of General Chemistry** J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section

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