

# percent composition

Positive... Chemistry!

What is the percent composition by mass of ethane,  $C_2H_6$ ?

$$\% \text{ comp. C} = \frac{2(12.01 \text{ g/mol})}{2(12.01 \text{ g/mol}) + 6(1.008 \text{ g/mol})} = \frac{24.02 \text{ g/mol}}{30.068 \text{ g/mol}} = 0.7987 = \boxed{79.87\% \text{ C}}$$

$$\% \text{ comp. H} = \frac{6(1.008 \text{ g/mol})}{2(12.01 \text{ g/mol}) + 6(1.008 \text{ g/mol})} = \frac{6.048 \text{ g/mol}}{30.068 \text{ g/mol}} = 0.2011 = \boxed{20.11\% \text{ H}}$$

shortcut for  
final element

$$\% \text{ comp. H} = 100\% - 79.89\% = \boxed{20.11\% \text{ H}}$$

$$\% = \frac{\text{part}}{\text{whole}}$$

$$\% \text{ composition X} = \frac{\text{molar mass X}}{\text{molar mass compound}}$$



# Percentage Composition Chemistry Answerkey

**Karin Nielsen-Saines**



## **Percentage Composition Chemistry Answerkey:**

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