CHEMISTRY: PERIODIC LAW "SCRAMBLE" LAB

WHAT TO TURN IN:	Unscrambled table	Questions #1-6

Introduction

The present organization of the elements is a product of the first periodic table published by Dmitri Mendeleev in 1869. The basis of his arrangement was the atomic masses of the elements. This approach proved incorrect as it would have placed some elements in a family with dissimilar properties. Henry Moseley rearranged the table on the basis of the atomic numbers of the elements. In accordance with Moseley's revision, the periodic law states the properties of the elements are periodic functions of their atomic numbers.

Each of the known elements has its own set of properties. These range from solid to gas, lustrous to dull, low to high melting points, various colors, etc. The elements are arranged within the periodic table into groups or families (vertical columns) and periods or rows (horizontal rows). This arrangement reflects the repeating nature of the properties of the elements.

In this experiment, you will use your knowledge of periodic properties and a list of clues to correctly arrange the elements from a scrambled periodic table.

Objectives

- To arrange the elements in Groups IA-VIIIA (1, 2, 13-18) according to a list of clues and your knowledge of periodic properties
- To explain the trends of properties in groups and periods

Equipment: scissors, glue stick, blank table, scrambled table of elements

Procedure

1. Obtain a "scrambled" table. Each block on the table represents a different element from Groups IA-VIIIA. (1-18)

2. Cut out blocks A-Z. Use the clues given at the end of the procedure and arrange the elements in their proper order the blank table provided by your teacher. When you are sure of a correct placement, glue the square in place. Elements A-Z will fill the first four rows.

3. Cut out the remaining 16 blocks (these have a * instead of a code letter). Use the information provided in each block and your knowledge of periodic properties to arrange these elements in their proper position on the blank table. Glue the blocks in place in rows five and six. THESE WILL BE AN EXTRA OPTIONAL GRADE AND WILL NOT AFFECT YOUR LAB SCORE IF YOU GET ANY INCORRECT OR DO NOT DO THEM.

4. Some information is missing from each block. Predict the values for the missing items from the location of the element on the periodic table. Place your predictions on the table. (You may use a periodic table *only* to determine the symbol for each element.) You must include the element symbol, atomic number, and oxidation number (charge):

GROUP NUMBER		GENERAL CHARGE (OXIDATION NUMBER) OF ION
Group IA	(1)	1+
Group IIA	(2)	2+
Group IIIA	(13)	3+
Group IVA	(14)	varies (mixed)
Group VA	(15)	3-
Group VIA	(16)	2-
Group VIIA	(17)	1-
Group VIIIA/0	(18)	0 (none)

Periodic Law Scramble Lab Answers

Gale Research Company

Periodic Law Scramble Lab Answers:

Federal Register ,1957 **Bulletin of the Atomic Scientists** ,1990-04 The Bulletin of the Atomic Scientists is the premier public resource on scientific and technological developments that impact global security Founded by Manhattan Project Scientists the Bulletin s iconic Doomsday Clock stimulates solutions for a safer world **Iron Age** ,1956 **Chief Executive** ,1993 **Motion Picture Story Magazine** ,1948 *Twentieth-century Literary Criticism* Gale Research Company,1978 Excerpts from criticism of the works of novelists poets playwrights and other creative writers 1900 1960

Los Angeles Magazine ,2002-09 Los Angeles magazine is a regional magazine of national stature Our combination of award winning feature writing investigative reporting service journalism and design covers the people lifestyle culture entertainment fashion art and architecture and news that define Southern California Started in the spring of 1961 Los Angeles magazine has been addressing the needs and interests of our region for 48 years The magazine continues to be the definitive resource for an affluent population that is intensely interested in a lifestyle that is uniquely Southern Californian

The Development of the Periodic Law Francis Preston Venable,1896 The Periodic Law Albert Edward Garrett,1909 *New Ideas in Chemistry from Fresh Energy for the Periodic Law* Henry Bent,2006-09-05 New Ideas calls to mind Aristotle s synopsis of the Iliad and the Odyssey Woman abducted Long war One guy has a hard time getting home End of story The rest is episodes Similarly here Chemical capture of the Left Step Periodic Table One element finds a new home The noblest of the noble gases is not a Noble Gas End of story The rest is novel consequences of the Noble Gas Conclusion Among them overlooked Rules of Triads Block Sizes and Full Shells overlooked block to block trends and a correspondence between elements ordinal numbers in their Groups and orbital s radial quantum numbers and recognition that Pauli s explanation of Periodicity s magic numbers 2 8 18 got the right answer the Pauli Exclusion Principle for the wrong reason New Ideas ends with suggestions for streamlining the teaching of the mole concept chemical bonding and thermodynamics in order to provide room in the chemistry curriculum for a more thorough treatment of Periodic System Systematics *On the Discovery of the Periodic Law* John A. R. Newlands,1884 By the English chemist whose work on the atomic weights of the elements anticipated the periodic table of Mendeleev and who predicted the element germanium before its discovery by the latter **The Development of the Periodic Law** Francis Preston Venable,2013-09 This historic book may have numerous typos and missing text Purchasers can usually download a free scanned copy of the original book without typos from the publisher Not indexed Not illustrated 1896 edition Excerpt it is obvious that Prout s law or some modification of it such as was many years ago suggested by Dumas must be true the atomic weights of all the other so called elements must be multiples of that of hydrogen or multiples of that fraction of the hydrogen atom which may result from the dissociation of this body itself If such fraction be very small as compared with the effect of the inevitable errors of experiment the experimental vertification or refutation of the law will prove impossible but if it be considerable as for instance onehalf of the commonly

known hydrogen atom or onefourth as assumed by Dumas the question admits of practical examination The author further questioned the justice of the view taken by Stas of his results that Prout s law is disproved by them or is not supported by them The careful work of Stas and others only proves by close agreement of the results that fortuitous errors have been reduced within narrow limits It does not prove that all sources of constant error have been avoided and indeed this can never be absolutely proved as we never can be sure that our knowledge of the substances we are dealing with is complete He added that of course one distinct exception to the assumed law would disprove it if that exception were itself fully proved but this is not the case Out of the eighteen best known atomic weights ten approximate to integers within a range of variation less than one tenth of a unit The degree of probability that this is purely accidental is found to be only equal to 1 1097 8 This seems to illustrate the point that not only is Prout s law not as yet absolutely over turned but that a heavy and **The Periodic Law** David Johnson,1984 *Mendeleev on the Periodic Law* Dmitri Ivanovich Mendeleev,2013-04-25 By the dawn of the nineteenth century elements had been defined as basic building blocks of nature resistant to decomposition by chemical means In 1869 the Russian chemist Dmitri Ivanovich Mendeleev organized the discord of the elements into the periodic table assigning each element to a row with each row corresponding to an elemental category The underlying order of matter hitherto only dimly perceived was suddenly clearly revealed This is the first English language collection of Mendeleev s most important writings on the periodic law Thirteen papers and essays divided into three groups reflect the period corresponding to the initial establishment of the periodic law three papers 1869 71 a period of priority disputes and experimental confirmations five papers 1871 86 and a final period of general acceptance for the law and increasing international recognition for Mendeleev five papers 1887 1905 A single easily accessible source for Mendeleev s principle papers this volume offers a history of the development of the periodic law written by the law s own founder *The Developement of the Periodic Law* Francis Preston Venable,1974 *Quiz Yourself Clever! The Periodic Table* DK,2024-10-03 Get to grips with the periodic table with this fact packed quiz book Learn all the ingredients that make up the universe in this quiz book for children aged 9 that breaks down the periodic table Quiz Yourself Clever The Periodic Table goes through all 118 elements in an accessible and kid friendly way showing the pure form of every element through an eye catching image surrounded by quiz style questions This periodic table quiz book for children offers Fast and fun learning through the use of detailed images and interesting facts Hundreds of quick fire questions for children to quiz themselves and challenge friends and family Information that has been endorsed and authenticated by experts Vivid pictures and fast facts will allow children to gather a wealth of knowledge in a fun way Turn the page to find the answers contained in a handy dataset from its atomic number and mass to its melting point and when it was discovered Images of how it appears in nature and how we use it give a fuller understanding of each element More in the series If you enjoyed Quiz Yourself Clever The Periodic Table then why not test yourself other quiz books to boost your knowledge like Quiz Yourself Clever Rocks Minerals to get to grips with the

fascinating rock and mineral forms all over the world or Animals of the World to learn about the most fascinating wildlife on our planet *On the Discovery of the Periodic Law, and on Relations Among the Atomic Weights* John A. R. Newlands,1884

The Basics of the Periodic Table Leon Gray,2013-12-15 Provides basic information on the periodic table Includes biographical information on Dmitri Mendeleev color photographs and diagrams sidebars a glossary and further reading sources **On the Discovery of the Periodic Law and on Relations Among the Atomic Weights** John A. R. Newlands,1844 Energy Wave of the Periodic Table William C. Scott,2007 Definitive Periodic Law is revealed in arrangement of new Periodic Table repeating sequential numbers of protons discovered in Groups 1 through 18 in the elements of the new ENERGY WAVE of the Periodic Table The elements in the new ENERGY WAVE of the Periodic Table are given in the ground state which is one electron for each proton This arrangement provided a unique opportunity to observe the nucleus of the elements By incorporating the sequential numbers of protons underlying the Energy Levels K L M N O P Q in shell blocks s p d and f of the Group elements it revealed what had been hidden and veiled in the complexity of electron configurations Sequential numbers of protons are observed to repeat in the Group elements from period to period This is the true revealed energy force creating the similar physical and chemical properties of Groups 1 through 18 from period to period in the Periodic Table The ENERGY WAVE of the Periodic Table had revealed Definitive Periodic Law Definitive Periodic Law is the number of protons underlying the Energy Levels K L M N O P Q in the nucleus of the Elements These sequential numbers of protons repeat in shell blocks s p d and f forming groups that have similar physical and chemical properties from period to period These sequential numbers of protons are the cornerstones of the nucleus and provide the atomic orbitals of the electrons the foundation for their spatial relationship to the nucleus as described by the azimuthal angular and magnetic numbers of quantum chemistry These sequential numbers of protons are very important as they reveal new explanation to chemical bond angles and the molecular geometry and structure of molecules

Periodic Law Scramble Lab Answers Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the ability of words has become more evident than ever. They have the ability to inspire, provoke, and ignite change. Such could be the essence of the book **Periodic Law Scramble Lab Answers**, a literary masterpiece that delves deep into the significance of words and their effect on our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book is key themes, examine its writing style, and analyze its overall effect on readers.

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